

Name: \_\_\_\_\_ per: \_\_\_\_\_

### Worksheet - Reaction Rates

Use this reaction for the questions below:  $C_6H_{12}O_6 (s) + 6 O_2(g) \rightarrow 6 H_2O (g) + 6 CO_2 (g)$

- What happens to the concentrations of:
  - $C_6H_{12}O_6$  &  $O_2$  as the reaction proceeds  $\rightarrow$ ?
  - $H_2O$  +  $CO_2$  as the reaction proceeds  $\rightarrow$ ?
- According to the collision theory, what 3 circumstances are needed for  $C_6H_{12}O_6$  &  $O_2$  to react?
- What is the activation energy for a chemical reaction?
- Use the equation & the collision theory to explain:



| Change in condition:                                                    | Does this <u>increase</u> or <u>decrease</u> the rate of reaction? | Explain why                                                   |
|-------------------------------------------------------------------------|--------------------------------------------------------------------|---------------------------------------------------------------|
| a. <u>Increasing the temperature</u>                                    | Ex: Increase (springs up)                                          | Ex: Molecules move faster & collide more = increased rxn rate |
| b. <u>Increasing the concentration</u> of $C_6H_{12}O_6$                |                                                                    |                                                               |
| c. <u>Decreasing the concentration</u> of $O_2$                         |                                                                    |                                                               |
| d. <u>Increase the surface area</u> by chewing up food in your mouth.   |                                                                    |                                                               |
| e. <u>Decreasing the temperature</u>                                    |                                                                    |                                                               |
| f. <u>Increasing the pressure</u> in the container                      |                                                                    |                                                               |
| g. <u>Decreasing the concentration</u> of $H_2O$                        |                                                                    |                                                               |
| h. <u>Increasing the volume</u> of the container the reaction occurs in |                                                                    |                                                               |
| i. <u>Increasing the concentration</u> of $CO_2$                        |                                                                    |                                                               |
| j. <u>Using a catalyst</u> (like salivary amylase)                      |                                                                    |                                                               |

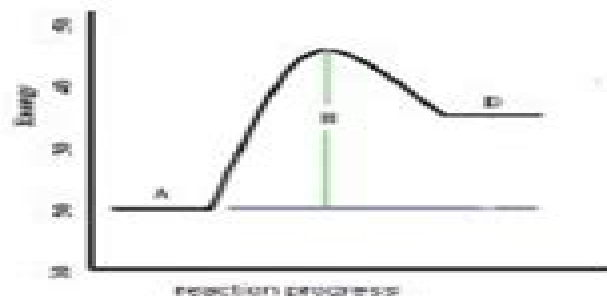
- On the accompanying energy diagram, label the following terms:

a. reactants    b. products    c. activation energy

- On the graph to the right, draw and label what this diagram would look like if a catalyst was added to the reaction.

- Graph reading

- How much energy (kJ) do the reactants have?
- How much energy (kJ) do the products have?
- How much energy (kJ) is required to activate this un-catalyzed reaction?
- Is this reaction endothermic or exothermic? How do you know?



# Reaction Rates Chapter Worksheet Page 99answers

**G Thomas**



**Reaction Rates Chapter Worksheet Page 99answers:**

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