

- Q6 When water was added to the anhydrate of $\text{CuSO}_4(x)$, did the reaction proceed in the direction of the reactants or products?

B.2 Reaction of Copper(II) Ion and Hydroxide

	Test tube 1	Test tube 2
Initial color of CuCl_2 solution (1)	light blue	light blue
Initial appearance $\text{Cu(OH)}_2(x)$ (2)	clear	clear
Change in appearance (3)	cloudy	clear
Change in appearance (4)	cloudy	clear

- Q7 In test tube 1, did the addition of OH^- cause the reaction to go in the direction of the reactants or products? Explain.
- Q8 In test tube 2, did the addition of H_3O^+ cause the reaction to go in the direction of the reactants or products? Explain.

C. Changing Equilibrium Conditions: Le Châtelier's Principle

C.1 Effect of Changing Reactant Concentration

Test tube	Color	Is color deeper or lighter red compared to the reference?	$[\text{FeSCN}^{2+}]$ increases or decreases?
1. Reference	red-orange	deeper/lighter	increases/decreases
2. Fe^{3+} added	dark red	deeper	increased
3. SCN^- added	blood red	deeper	increased
4. Fe^{3+} removed	orange	lighter	decreased

Reaction Rates And Equilibrium Lab Answers

R Bogdan



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