

# Reaction Rates and Chemical Equilibrium



## Glossary

| TERM                 | DEFINITION                                                                                                                                                    |
|----------------------|---------------------------------------------------------------------------------------------------------------------------------------------------------------|
| activation energy    | the minimum amount of energy needed to start a chemical reaction                                                                                              |
| catalyst             | a substance that lowers the activation energy of a reaction but is otherwise not part of the reaction                                                         |
| chemical equation    | the symbolic representation of a chemical reaction can be written as                                                                                          |
| chemical equilibrium | a state of a reversible reaction in which the forward and reverse reactions occur at the same rate and the quantities of reactants and products do not change |
| chemical reaction    | a process that involves the rearrangement of atoms, molecules, or both to form new substances                                                                 |
| concentration        | the amount of solute in a given amount of solvent or solution                                                                                                 |
| endothermic reaction | a chemical reaction that absorbs energy                                                                                                                       |
| energy               | the property of a body or a system that enables it to do work or cause a change                                                                               |
| exothermic reaction  | <b>a chemical reaction that releases energy</b>                                                                                                               |
| product              | <b>a substance formed in a chemical reaction</b>                                                                                                              |
| reactant             | a substance that is present at the beginning of a chemical reaction and is transformed by the reaction                                                        |
| reaction rate        | the rate at which reactants are used up or products are formed in a chemical reaction                                                                         |
| reversible reaction  | a chemical reaction that occurs simultaneously in the forward and reverse directions, converting reactants to products and products to reactants              |

## Objective

In this lesson, you will predict how altering the environment in which a chemical reaction takes place changes the rate of the reaction and the amounts of products.

## Reaction Rates

In a chemical reaction, atoms rearrange themselves by breaking and forming new bonds. The atoms of the reactants rearrange to form the products, which are made of the same atoms as the reactants. The time taken for reactants to convert to products is the reaction rate. As a reaction progresses, the amount of reactants decreases and the amount of products increases.

# Reaction Chemistry Rates And Equilibrium Guided Answers

**Disha Experts**



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