

Chapter 9

9.1 Compounds and Chemical Change

- **Atom**- the smallest unit of an element that can exist alone or in combination with other elements.
- **Compounds**- formed when atoms are held together by an attractive force called a **chemical bond**
- **Molecule**- The smallest particle of a compound or a gaseous element, that can exist and still remain the characteristic chemical properties of a substance
- **Chemical reaction**- a change in matter where **different chemical substances are created by breaking and/or forming chemical bonds.**
 - Examples: Photosynthesis
- Energy is *absorbed* -> produce new chemical substances -> more internal potential energy
- Energy is *released* when new chemical substances are produced -> less internal potential energy
- **Chemical Equation**: a way of describing what happens during a chemical reaction
 - Example: Photosynthesis equation-
Energy/Sunlight + carbon dioxide molecules + water molecules -> plant material molecules + oxygen molecules

9.2 Valence Electrons and Ions

- the number of electrons in the **outermost orbital** that **usually determines the chemical properties of an atom**
- **Valence electron**:- the number of electrons in the outermost orbital that usually determines the properties of an atom
- **Octet Rule**- after the element helium, atoms attempt to acquire an outer orbital with eight electrons through chemical reactions
-

9.3 Chemical Bonds

- **Chemical bond**-an attractive force that holds atoms together in a compound.
- Three Classes of Chemical Bonds-
 - Ionic Bonds
 - Covalent Bonds
 - Metallic Bonds
- **Ionic Bonds**- type of linkage formed from the electrostatic attraction between oppositely charged ions in a chemical compound
- **Covalent bonds**- result when atoms achieve the noble gas electron structure by *sharing* electrons. They generally formed between the nonmetallic elements on the right side of the periodic table
- **Metallic bonds**- formed in solid metals such as iron, copper, and the other metallic elements that make up about 80 percent of all the elements.
- **Heat Formation**- energy that is released during an ionic bond formation
- **Formula**- describes what elements are in the compound and in what proportions
- **Single bond**- covalent bond
- **Double bond**- formed by two pairs of electrons that are shared by two atoms

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Table of Contents Physical Science Chapter Vocabulary

1. Understanding the eBook Physical Science Chapter Vocabulary
 - The Rise of Digital Reading Physical Science Chapter Vocabulary
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter Vocabulary
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter Vocabulary
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter Vocabulary
 - Personalized Recommendations
 - Physical Science Chapter Vocabulary User Reviews and Ratings
 - Physical Science Chapter Vocabulary and Bestseller Lists

5. Accessing Physical Science Chapter Vocabulary Free and Paid eBooks
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 - Physical Science Chapter Vocabulary Budget-Friendly Options
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 - Physical Science Chapter Vocabulary Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter Vocabulary
 - Highlighting and Note-Taking Physical Science Chapter Vocabulary
 - Interactive Elements Physical Science Chapter Vocabulary
8. Staying Engaged with Physical Science Chapter Vocabulary
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter Vocabulary
9. Balancing eBooks and Physical Books Physical Science Chapter Vocabulary
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter Vocabulary
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Chapter Vocabulary
 - Setting Reading Goals Physical Science Chapter Vocabulary
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Chapter Vocabulary
 - Fact-Checking eBook Content of Physical Science Chapter Vocabulary
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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