



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use blue or black ink for writing. Do not use red ink.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹ and 1 mol of 1-hexanol.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B? or C? and list one of the reasons. (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and is the prime reason of their volatility. (2)
- 3.1.4 What factor of the compounds A and B is responsible for the difference in their boiling points? (2)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 Which compound (A or B) will have the higher surface tension? (2)

Compound C (M_r 102) will show in the table, has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 16,2 °C.

- 3.1.8 Explain why the boiling point of compound C is 16,2 °C (lower than that of compound A and B) by referring to the ΔH_{vap} and intermolecular or intramolecular forces involved. (2)

3.2 Substances D and E are shown in their condensed state. The diagram below shows the molecular structure of the substances.



- 3.2.1 Is substance D or E or both, soluble in water? (2)
- 3.2.2 Explain the process in QUESTION 3.1.1 about the volatility in the intermolecular forces involved. (2)

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A Gutmann



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