

Chapter 5 Study Guide- Physical Science
Test Monday, April 28th, 2014

1. What is a chemical bond? **force of attraction that holds elements together in order to form a unit**
2. Define chemical structure. **the arrangement of atoms in a substance.**
3. What is an ionic bond? **a chemical bond in which one atom loses an electron to form a positive ion and the other atom gains an electron to form a negative ion**
4. An ionic bond is between a **METAL** and a **NONMETAL** .
5. What is a covalent bond? **a chemical bond that involves sharing a pair of electrons between atoms in a molecule**
6. A covalent bond is between a **nonmetal** and a **nonmetal** .
7. What is a metallic bond? **a bond formed by the attraction between positively charged metal ions and the electrons around them.**
8. What are the polyatomic ions? **an ion made of two or more atoms.**
9. What is an ION? **a charged atom**
10. What is a CATION? **A positively charged atom**
11. What is an ANION? **A negatively charged atom**
12. Which part of an ionic compound should be written first? **Metal**

Name the following ionic compounds:

13. FeI₂- **Iron Iodide**
14. MnF₃- **Manganese Fluoride**
15. CrCl₂- **Chromium Chloride**
16. CuS- **Copper Sulfide**

Name the following covalent compounds:

17. As₂O₅- **diarsenic pentaoxide**
18. SiI₄- **silicon tetraiodide**
19. P₄S₃- **terraphosphorus trisulfide**
20. P₄O₁₀- **tetraphosphorus decoxide**

Write the equation for the following ionic compounds:

21. Lithium Oxide- **Li₂O**
22. Magnesium Nitride- **Mg₃N₂**
23. Sodium Phosphide- **Na₃P**
24. Calcium Bromide -**CaBr₂**

Define/Explain the following

25. What is a subscript? **number written below, tells you how many of each element is present in compound/molecule**
26. What is a superscript? **number written above the element symbol, this is how the oxidation number (charge) is displayed.**

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Michael Seilmaier



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