

Periodic Trends Worksheet

Using the periodic table and your knowledge of periodic trends, answer the following questions:

1 Circle the best answer that has the:

- | | | | | | | |
|---|----------------------------|----|----|----|----|----|
| a | smallest electron affinity | N | P | As | Sb | Bi |
| b | largest electronegativity | Ga | Ge | As | Sc | |
| c | largest ionization energy | O | S | Sc | Te | Po |
| d | lowest ionization energy | O | S | Se | Te | Po |
| e | smallest atomic radius | F | Cl | Br | I | At |
| f | largest atomic radius | Li | Na | K | Rb | Cs |

2 Low ionization energy is a characteristic of a _____

3 What does atomic radius measure? _____

4 Describe the change in the atomic size across a period and down a group.

5 Which element has the largest atomic radius? _____

6 Explain the main difference between metals and nonmetals.

7 Determine if the listed pairs of elements have similar or different properties.

- | | |
|------------------------|----------------------|
| a Potassium & Rubidium | b Calcium & Barium |
| c Sodium & Chlorine | d Helium & Krypton |
| e Phosphorus & Oxygen | f Lithium & Fluorine |

8 Which element has the largest electronegativity? _____

9 Which group (name) would have the lowest ionization energies?

10 Which group (name) has the highest ionization energies?

Periodic Trends And Ions Review Answer Sheet

Kaplan Test Prep



Periodic Trends And Ions Review Answer Sheet:

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