

Boyle's Law

- The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
- math
- math

Kinetic Theory and Boyle's Law

- Pressure of a gas is caused by the collisions of the gas molecules impinging the container.
- If the gas is compressible to small the volume it has large number molecules the pressure is very large.
- Force moving rapidly over small on the walls of the container.
- Speed of the gas will increase.

- Ex 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. If the gas is released, it maintains a volume of 900 ml. above the pressure inside of 800 Pa. What is the volume of the balloon at this altitude?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 101.3 \text{ kPa} \quad P_2 = 80 \text{ kPa} \quad V_1 = 800 \text{ ml} \quad V_2 = ?$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{101.3 \text{ kPa} \times 800 \text{ ml}}{80 \text{ kPa}} = 1013 \text{ ml}$$

- Ex 2. A gas at pressure of 101.3 kPa and 400 ml. of a certain gas has a volume of 900 ml. If the pressure is increased until the volume is 800 ml. what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 101.3 \text{ kPa} \quad V_1 = 400 \text{ ml} \quad V_2 = 800 \text{ ml} \quad P_2 = ?$$

$$P_2 = \frac{P_1 V_1}{V_2} = \frac{101.3 \text{ kPa} \times 400 \text{ ml}}{800 \text{ ml}} = 50.65 \text{ kPa}$$

Note Taking Guide Episode 902 Chemistry

Padhraic Smyth



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