

What are Acids, Bases, and Salts?

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form H^+ ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form OH^- ions.



- Salts are compounds that dissociate in aqueous solution releasing neither H^+ ions nor OH^- ions.



Using the **Arrhenius definition**, classify the following examples as acids, bases, or salts:

HBr	<u>acid</u>	KCl	<u>salt</u>
Mg(OH) ₂	<u>base</u>	H ₃ PO ₄	<u>acid</u>
HCl	<u>acid</u>	HClO	<u>acid</u>
KNO ₃	<u>salt</u>	Al(OH) ₃	<u>base</u>
HFO ₄	<u>acid</u>	KC ₂ H ₃ O ₂	<u>salt</u>
Ba(OH) ₂	<u>base</u>	NaCl	<u>salt</u>

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Enjoying the Tune of Expression: An Psychological Symphony within **Note Taking Worksheet Acids Bases And Salts**

In some sort of eaten by monitors and the ceaseless chatter of instantaneous transmission, the melodic beauty and mental symphony produced by the written word frequently disappear in to the backdrop, eclipsed by the relentless noise and disturbances that permeate our lives. But, located within the pages of **Note Taking Worksheet Acids Bases And Salts** a wonderful fictional prize filled with natural thoughts, lies an immersive symphony waiting to be embraced. Constructed by a masterful musician of language, that interesting masterpiece conducts viewers on a psychological trip, well unraveling the hidden melodies and profound affect resonating within each carefully constructed phrase. Within the depths of this moving evaluation, we can explore the book is central harmonies, analyze its enthralling publishing type, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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