

Calculus-Based Properties (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \times 5.56 \text{ mol/kg} = 10.3^\circ C \quad \Delta T_b = K_b \cdot m = 0.512^\circ C/m \times 5.56 \text{ mol/kg}$$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of NaCl and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$, $K_f(\text{NaCl}) = 3.72^\circ C/m$, $K_b(\text{NaCl}) = 1.02^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{0.100 \text{ kg}} = 1.71 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 3.72^\circ C/m \times 1.71 \text{ mol/kg} = 6.36^\circ C \quad \Delta T_b = K_b \cdot m = 1.02^\circ C/m \times 1.71 \text{ mol/kg}$$

3. How many grams of $C_6H_{12}O_6$ must be added to 100 g of H_2O to raise the boiling-point of the solution to $100.04^\circ C$? Assume that the solute does not dissociate. Equations: $K_b = 0.512^\circ C/m$.

$$\Delta T_b = K_b \cdot m \quad 0.04^\circ C = 0.512^\circ C/m \times m \quad m = 0.0781 \text{ mol/kg}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.0781 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$x = 0.00781 \text{ mol} \times 180 \text{ g/mol} = 1.41 \text{ g}$$

$$m = 1.41 \text{ g}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.0781 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

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4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 100 g of water is prepared. The freezing-point of this solution is $-0.54^\circ C$. The normal freezing-point of pure water is $0^\circ C$ and the freezing-point depression constant for water is $1.86^\circ C/m$. Determine the molar mass of the compound. Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\Delta T_f = K_f \cdot m \quad 0.54^\circ C = 1.86^\circ C/m \times m \quad m = 0.290 \text{ mol/kg}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.290 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$x = 0.0290 \text{ mol} \times 180 \text{ g/mol} = 5.22 \text{ g}$$

$$m = 5.22 \text{ g}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.290 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$x = 0.0290 \text{ mol} \times 180 \text{ g/mol} = 5.22 \text{ g}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = 0.290 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m \quad 0.54^\circ C = 1.86^\circ C/m \times m \quad m = 0.290 \text{ mol/kg}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.290 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}} \quad x = 0.0290 \text{ mol} \times 180 \text{ g/mol} = 5.22 \text{ g}$$

Rosengarten Colligative Properties Homework Answers

Tobias Bleicker



Rosengarten Colligative Properties Homework Answers:

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