

CHAPTER 13

Rates of Reaction

■ SOLUTIONS TO EXERCISES

Acetaminophen is a painkiller that gives analgesic results in the treatment of headaches, whereas aspirin can relieve pain, but the two agents perform in contrasting ways. The high amount of the aspirin in the system indicates aspirin's effectiveness in eliminating pain, whereas acetaminophen's low level indicates only partial relief has been achieved.

110. Data at intervals of 10 s: $x = 10, 20, 30, 40$. Data values: $0.0010, 0.0015, 0.0020, 0.0025$. Divide each rate by the initial concentration of the corresponding reactant in each equation:

$$\text{a. } \frac{0.0010}{0.01} = 0.10, \quad \frac{0.0015}{0.01} = 0.15, \quad \frac{0.0020}{0.01} = 0.20, \quad \frac{0.0025}{0.01} = 0.25$$

111. Data: $x = \frac{0.01}{2} = 0.005$, $\frac{0.01}{3} = 0.0033$, $\frac{0.01}{4} = 0.0025$, $\frac{0.01}{5} = 0.002$, $\frac{0.01}{6} = 0.0017$

112. The data table appears to follow the pattern of Table 13.1. The concentration of O_2 is the same, whereas the volume of H_2 is

113. Hydrogen peroxide decomposes to water and the corresponding substance that is evolved. Thus, the amount of evolved oxygen that corresponds to each

time is shown in

The concentration of the reacting substance, represented in the table, has not been determined. Based on what Table 13.1 shows:

$$k = \frac{\text{rate}}{[\text{reactant}]^n} = \frac{0.0001 \text{ mol/L} \cdot \text{s}}{(0.001 \text{ mol/L})^2} = 10 \text{ L/mol} \cdot \text{s} \text{ (constant)}$$

114. a. For $[\text{H}_2\text{O}_2]$ and $[\text{H}^+]$ at 0.01 mol/L, the reaction rate was the same, indicating that the concentration of H^+

$$\text{is } \frac{0.0001}{(0.01)^2} = 1 \text{ L/mol} \cdot \text{s} \text{ (constant) and } \frac{0.0001}{(0.01)^2} = 1 \text{ L/mol} \cdot \text{s} \text{ (constant)}$$

For a the following relationship exists:

$$\frac{0.0001}{(0.01)^2} = \frac{0.0001}{(0.01)^2} \text{ (constant)}$$

Thus:

$$[\text{H}_2\text{O}_2] = 0.01 \text{ mol/L} \text{ at } 0.0001 \text{ L/mol} \cdot \text{s} \text{ (constant) and } 0.0001 \text{ L/mol} \cdot \text{s} \text{ (constant)}$$

Reaction Rates Assessment Solutions Manual

AW Rasmussen



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