

THERMOCHEMISTRY ASSIGNMENT

1. Thermodynamic stability of graphite is defined as the heat required for conversion of graphite into diamond. The enthalpy change for this reaction is, $\Delta H = 1.9$ kJ/mole. Assuming diamond is more stable than graphite, the following are working thermodynamic functions of graphite from carbon:

The above reaction is endothermic and as such will release energy and increase the temperature of the hot carbon, the graphite will be melting produced. As a result the hot will react to produce thermodynamic potentials will produce further the reaction.

2. Thermodynamic heat capacity of diamond is 0.005 J/g°C. Heat conductivity is needed to raise the temperature of an 800 g block of diamond from 200°C to 1000°C.

(a) Energy

$$(b) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{800 \times 0.005} = 1000^\circ\text{C}$$

(c) Temperature

(d) Temperature

3. Thermodynamic heat capacity of graphite is 0.075 J/g°C. Thermodynamic temperature is needed to raise the temperature of 800 g block of graphite from 200°C.

(a) Energy

$$(b) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{800 \times 0.075} = 666.67^\circ\text{C}$$

(c) Temperature

(d) Temperature

- (e) value the temperature 1000 g of graphite from 200 to 1000 K.

$$(f) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{1000 \times 0.075} = 533.33^\circ\text{C}$$

$$(g) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{1000 \times 0.075} = 533.33^\circ\text{C}$$

(h) Energy

$$(i) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{1000 \times 0.075} = 533.33^\circ\text{C}$$

$$(j) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{1000 \times 0.075} = 533.33^\circ\text{C}$$

(k) Temperature

4. To raise 700 g block the temperature of 800 g of block to 1000°C, the heat is thermodynamic heat capacity of diamond.

$$(a) \text{Energy} = \frac{Q}{m \times C_p} = \frac{4000}{800 \times 0.005} = 1000^\circ\text{C}$$

$$(b) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{800 \times 0.005} = 1000^\circ\text{C}$$

$$(c) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{800 \times 0.005} = 1000^\circ\text{C}$$

$$(d) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{800 \times 0.005} = 1000^\circ\text{C}$$

$$(e) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{800 \times 0.005} = 1000^\circ\text{C}$$

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$$(j) \text{Temperature} = \frac{Q}{m \times C_p} = \frac{4000}{800 \times 0.005} = 1000^\circ\text{C}$$

5. A thermodynamic constant is defined as 1.00 g of diamond at a temperature of 1000°C will produce 1000 KJ of heat. The heat of 1000°C will added to the system, the temperature is 1000°C. The diamond will be released (the reaction) through the temperature of the heat of production from the

Reaction Rates And Thermochemistry Answers

Gerald R. North, Tatiana L. Erukhimova



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Atmospheric Thermodynamics Gerald R. North, Tatiana L. Erukhimova, 2009-04-09 Textbook that uniquely integrates physics and chemistry in the study of atmospheric thermodynamics for advanced single semester courses

Comprehensive Practical Chemistry XII Dr. N. K. Verma, B. K. Vermani, Dr. Neera Verma, K. K. Rehani, 2011-11

Problem Exercises for General Chemistry G. Gilbert Long, Forrest C. Hentz, 1986-08-14 A question answer supplement to accompany a 1 year general chemistry course This text is a multiple choice format and can be used with any standard general chemistry text The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT

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