

Key to Reaction Prediction -1

For each of the following eight reactions, in part (i) write a **BALANCED** equation and in part (ii) answer the question about the reaction. In part (i), coefficients should be in terms of **lowest whole numbers**. Assume that solutions are aqueous unless otherwise indicated. Represent substances in solutions as ions if the substances are extensively ionized. Omit formulas for any ions or molecules that are unchanged by the reaction.

- a) Solid Calcium Carbonate is strongly heated.



- ii. What is the oxidation state of the carbon and does it change?



- b) Hydrogen sulfide gas is bubbled into a solution of mercury (II) chloride.



- ii. What would you expect to observe in the solution during this reaction.

A black precipitate forms.

- c) Phosphorus(V) oxide powder is sprinkled over distilled water.



- ii. Estimate the pH of the reacted solution and give your reasoning.

pH would be less than 7 as a weak acid is formed

- d) A bar of zinc metal is immersed in a solution of copper(II) sulfate.



- ii. What is the oxidizing agent? Explain.

CuSO_4 is the oxidizing agent. Copper gains electrons.

- e) A small piece of sodium metal is added to distilled water.



- ii. A flame is seen above the water. Explain this secondary reaction.

Hydrogen gas is flammable and the reaction is exothermic enough to ignite the gas.

- f) Solid barium oxide is added to distilled water.



- ii. Estimate the pH of the reacted solution and give your reasoning.

The pH should be greater than 7 since a base is made.

- g) Solid ammonium carbonate is heated.



- ii. If one mole of reactant is used what is the volume of the product STP.

1 mole of reactant should yield three moles of product.

- h) Solutions of manganese(II) sulfate and ammonium sulfide are mixed.



- ii. Name the spectator ion(s).

Ammonium and sulfate are the spectator ions.

Reaction Prediction Practice Answer Key

JR Anderson



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Table of Contents Reaction Prediction Practice Answer Key

1. Understanding the eBook Reaction Prediction Practice Answer Key
 - The Rise of Digital Reading Reaction Prediction Practice Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Prediction Practice Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Prediction Practice Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Prediction Practice Answer Key
 - Personalized Recommendations
 - Reaction Prediction Practice Answer Key User Reviews and Ratings
 - Reaction Prediction Practice Answer Key and Bestseller Lists
5. Accessing Reaction Prediction Practice Answer Key Free and Paid eBooks
 - Reaction Prediction Practice Answer Key Public Domain eBooks
 - Reaction Prediction Practice Answer Key eBook Subscription Services
 - Reaction Prediction Practice Answer Key Budget-Friendly Options

6. Navigating Reaction Prediction Practice Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Reaction Prediction Practice Answer Key Compatibility with Devices
 - Reaction Prediction Practice Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Prediction Practice Answer Key
 - Highlighting and Note-Taking Reaction Prediction Practice Answer Key
 - Interactive Elements Reaction Prediction Practice Answer Key
8. Staying Engaged with Reaction Prediction Practice Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Prediction Practice Answer Key
9. Balancing eBooks and Physical Books Reaction Prediction Practice Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Prediction Practice Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Reaction Prediction Practice Answer Key
 - Setting Reading Goals Reaction Prediction Practice Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reaction Prediction Practice Answer Key
 - Fact-Checking eBook Content of Reaction Prediction Practice Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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