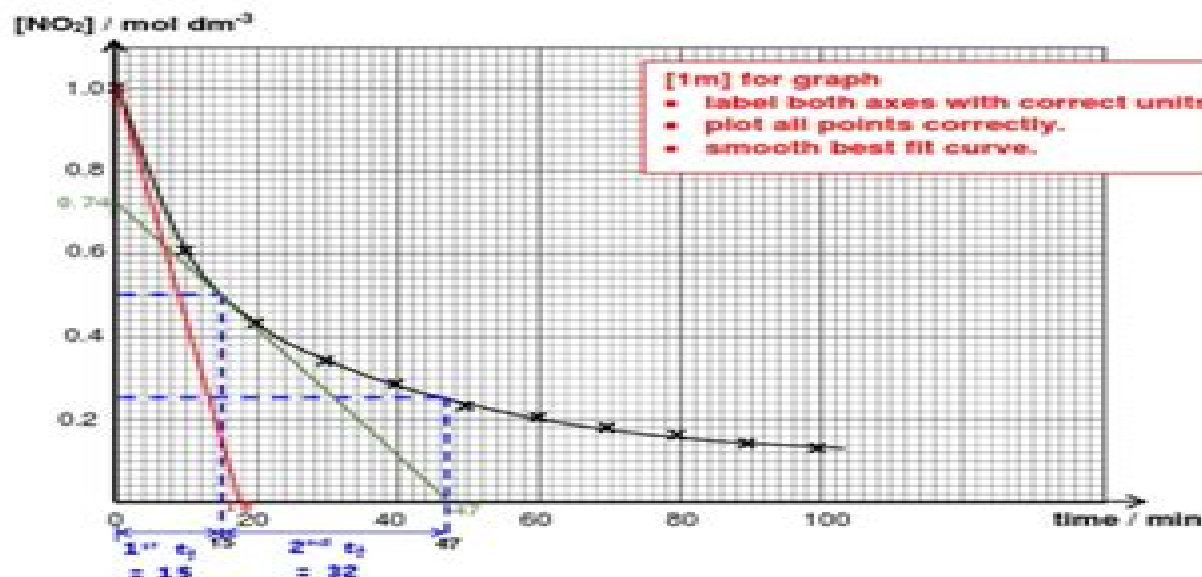


Topic 6: Reaction Kinetics (Answers)

1 (a)



[1m] Indicate and label values of two half-lives clearly on the graph using construction lines

Since half-lives are not constant, the reaction is not first order. [1m]

(b) (i) Initial rate = $\frac{1.0}{18} = 0.0556 \text{ mol dm}^{-3} \text{ min}^{-1}$ accept range of 0.0500 to 0.0625

[1m] tangent line drawn at $t = 0$ + value in 3sf + units

(ii) rate when $[\text{NO}_2] = 0.5 \text{ mol dm}^{-3} = \frac{0.74}{47} = 0.0157 \text{ mol dm}^{-3} \text{ min}^{-1}$

[1m] tangent line drawn at $[\text{NO}_2] = 0.5 \text{ mol dm}^{-3}$ + value in 3sf + units

(c)

$[\text{NO}_2] / \text{mol dm}^{-3}$	rate / $\text{mol dm}^{-3} \text{ min}^{-1}$
1.00	0.0556
0.50	0.0145

When $[\text{NO}_2]$ increases to 2 times its original value, the rate increases to ≈ 4 times its original value. } [1m]

Therefore, the order of reaction with respect to $[\text{NO}_2]$ is 2. [1m]

Reaction Energy Reaction Kinetics Answers

RC Schank

A decorative red curved shape, resembling a stylized arrow or a partial circle, is positioned to the right of the author's name.

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