Stoichiometry Using Molarity Practice Problems

Directions: Answer each question following a chemical equation

★ If 43.2 mL of 0.236 M NaOH reacts with 36.7 mL of H₂SO₄, what is the concentration of the H₂SO₄ solution?

★ If 36.7 mL of HCl solution is needed to react with 43.2 mL of a 0.236 M NaOH, what is the concentration of the HCl solution?

$$Al(OH)_3 + HCl \rightarrow AlCl_3 + H_2O$$

★ How many mL of 1.2 M HCl is needed to dissolve 5.8 g of Al(OH);?

$$H_2SO_4 + Na_2CO_3 \rightarrow Na_2SO_4 + H_2O + CO_2$$

★ Calculate the molarity of the H₂SO₄ solution if it takes 40.0 mL of H₂SO₄ to neutralize 0.364 g of Na₂CO₃.

$$Ca(OH)_2(s) + HCl(aq) \rightarrow CaCl_2(aq) + H_2O(1)$$

- ★ How many liters of 0.100 M HCl would be required to react completely with 5.00 grams of calcium hydroxide?
- ★ If I combined 15.0 grams of calcium hydroxide with 75.0 mL of 0.500 M HCl, how many grams of calcium chloride would be formed?

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WJ Hussar

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