



NOVEMBER EXAMINATION PAPER 2: CHEMISTRY

QUESTION 1

1.1 Which of the following is an example of a mixture?

- Element with the electron configurations: $1s^2, 2s^2, 2p^4$

- 1.1 The chemical formula of sodium sulfide

- | | |
|---|-------------------------|
| A | Fe_2S_3 |
| B | Fe_3S_4 |
| C | Fe_2S_2 |
| D | Fe_3S_2 |

- ☐ A Ion-dipole
- ☐ B Dipole-dipole
- ☐ C Van der Waals
- ☐ D Hydrogen bond

- 1.5 The particles of a gas ...

- A are tightly packed.
- B vibrate against each other.
- C move freely and rarely touch.
- D move over one another.

- 1.6 Element with the highest electronegativity?

- A. Hydrogen
 B. Helium
 C. Oxygen
 D. Fluorine

- 1.7 Within an atom of Group 16, it rises to become atom 26

- A mass number of the atom will increase
B atomic number of the atom will decrease
C number of electrons will decrease
D number of electrons will increase

- 1.18 An element with the following Aufbau diagram will be classified as a

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99	100
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- A noble gas
B halogen
C alkali earth-metal
D alkali metal

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SJ Ball



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