

Chemical Equilibrium Lab

Introduction:

Some chemical systems reach the mutually back-and-forth breaking-and-forming bonds simultaneously, resulting in a mixture of products and reactants. These reactions are called **reversible reactions**, and the reaction can proceed in either direction. Once the **equilibrium** is reached, the rates of the **forward** and **reverse** reactions are equal. **Le Chatelier's Principle** is the important concept used to explain how a system at equilibrium reacts to changes in its conditions. For instance, if a reaction system is at equilibrium, then changing the concentration of either the **reactants** or **products** will shift the equilibrium position to counteract the change. For example, if the concentration of a reactant is increased, the equilibrium will shift to the right to produce more products. Conversely, if the concentration of a product is increased, the equilibrium will shift to the left to produce more reactants. In this lab, you will be observing the changes in equilibrium when the amount of reactants, product, or temperature is changed. Follow the **Lab Procedure** for the following reactions. **Observe** the color changes that occur by the color changes in the solution, according to the color indicators in the equations below. For example, if the color changes to red, then reaction **1** is occurring, while in reaction **2**, the color change is to colorless.



Adding NaOH solution to $\text{Fe}^{3+}(\text{aq})$ and $\text{SCN}^{-}(\text{aq})$ will cause $\text{Fe}^{3+}(\text{aq})$ to react with $\text{OH}^{-}(\text{aq})$, forming the insoluble precipitate $\text{Fe}(\text{OH})_3$ in solution. This will produce more $\text{SCN}^{-}(\text{aq})$ in the solution, shifting the equilibrium to the right.

Procedure (This Group):

- | | |
|--|--|
| 1. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 2. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 3. Prepare 100 mL of 0.1 M NaOH solution. | 4. Prepare 100 mL of 0.1 M HCl solution. |
| 5. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 6. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 7. Prepare 100 mL of 0.1 M NaOH solution. | 8. Prepare 100 mL of 0.1 M HCl solution. |
| 9. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 10. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 11. Prepare 100 mL of 0.1 M NaOH solution. | 12. Prepare 100 mL of 0.1 M HCl solution. |
| 13. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 14. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 15. Prepare 100 mL of 0.1 M NaOH solution. | 16. Prepare 100 mL of 0.1 M HCl solution. |

Observing Chemical Equilibrium Lab 46

BM King

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

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Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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layers to the deeper atmosphere of the tropics Thermodynamic and kinematic consequences of these transfers of energy are extended to the geochemical and living worlds **Experiment Station Record** United States. Office of Experiment Stations,1913 **The Iron(III) Thiocyanate Reaction** Kevin C. de Berg,2019-11-13 This Brief presents an historical investigation into the reaction between ferric ions and thiocyanate ions which has been viewed in different ways throughout the last two centuries Historically the reaction was used in chemical analysis and to highlight the nature of chemical reactions the laws of chemistry models and theories of chemistry chemical nomenclature mathematics and data analysis and instrumentation which are important ingredients of what one might call the nature of chemistry Using the history of the iron III thiocyanate reaction as a basis the book s main objective is to explore how chemistry develops its own knowledge base how it assesses the reliability of that base and how some important tools of the trade have been brought to bear on a chemical reaction to achieve understanding a worthwhile goal of any historical investigation International Aerospace Abstracts ,1999 Observation, Prediction and Simulation of Phase Transitions in Complex Fluids Marc Baus,L.F Rull,Jean-Paul Ryckaert,2012-12-06 Observation Prediction and Simulation of Phase Transitions in Complex Fluids presents an overview of the phase transitions that occur in a variety of soft matter systems colloidal suspensions of spherical or rod like particles and their mixtures directed polymers and polymer blends colloid polymer mixtures and liquid forming mesogens This modern and fascinating branch of condensed matter physics is presented from three complementary viewpoints The first section written by experimentalists emphasises the observation of basic phenomena by light scattering for example The second section written by theoreticians focuses on the necessary theoretical tools density functional theory path integrals free energy expansions The third section is devoted to the results of modern simulation techniques Gibbs ensemble free energy calculations configurational bias Monte Carlo The interplay between the disciplines is clearly illustrated For all those interested in modern research in equilibrium statistical mechanics

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web shown above is the standard gm wiring diagram for this plug separate black wire with green connector is for the illuminated gear indicator connect this to dash lights or a switched 12 volt source steering column and ignition switch wiring conversion from gm wire color codes for ford and chrysler applications

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