Volumetric Analysis Calculations

 $30.0 cm^3$ of $0.100 moldm^{-3} NaOH$ reacted completely with $25.0 cm^3$ of H_2SO_4 in a titration.

Calculate the concentration of H_2SO_4 in $moldm^{-3}$?

Step 2: Find no' of moles in NaOH

$$2NaOH_{(aq)} + H_2SO_{4(aq)} \rightarrow Na_2SO_{4(aq)} + 2H_2O_{(l)}$$

 $V = 0.1 mold m^{-3}$ $V = 25 cm^3$

Using
$$n = c \times v$$

$$n(NaOH) = 0.1 \times \frac{30}{1000}$$

$$= 0.003 \, mol$$



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Yu, Holly, Breivold, Scott

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