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KATHLEEN SHAGINAW



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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web a write a balanced equation for the reaction Al + O₂ → Al₂O₃ 4 Al + 3 O₂ → 2 Al₂O₃ b determine the theoretical yield of Al₂O₃ 160.0 g Al x 1 mol Al / 27.0 g Al x 2 mol Al₂O₃ / 3 mol Al = 101.964 g Al₂O₃ 302.3 g Al₂O₃ 26.982 g Al x 4 mol Al / 1 mol Al₂O₃ = 107.936 g Al₂O₃ c determine the percent yield yield actual x 100 = 260.0 g / 260.0 g x 100 = 86.01% theoretical 302.3 g

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