



PA-DSS Implementation Guide for Sage MAS 90 and 200 ERP Credit Card Processing

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Sage Mas 5sdk Guide Api

Padhraic Smyth



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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$ total ionic $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$ net ionic $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water
 $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ the spectator ions in $\text{KOH}(\text{aq})$ and $\text{HNO}_3(\text{aq})$ are $\text{K}^+(\text{aq})$ and $\text{NO}_3^-(\text{aq})$ how to identify and list the

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 $\text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{KCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{KNO}_3(\text{aq})$ complete ionic equation an equation written with all soluble

strong electrolytes shown as ions

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web the Ag^+ and Cl^- ions become AgCl s but the Na^+ ions and the NO_3^- ions stay as Na^+ ions and NO_3^- ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{Cl}^- + \text{Cu}^{2+} + 2\text{aq} \rightarrow \text{CuCl}_2 \text{ s}$ $2\text{Cl}^- + \text{Cu}^{2+} + 2\text{aq} \rightarrow \text{CuCl}_2 \text{ s}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length

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