

AP Chemistry: A Sample Formal Laboratory Report

This paper is designed to help you prepare a chemistry lab report. Keep it in your chemistry notebook. All chemistry lab reports must be written neatly and well organized to receive full credit. Lab reports may be written or typed. It is highly recommended that you use graph or engineering bond paper for written reports.

Laboratory #7: Quantitative Determination of an Empirical Formula

I. Hypothesis: If nitric acid is poured onto tin, a tin oxide will be produced. If we know the initial mass of the tin metal and the mass of the final product, we can determine the empirical formula of the tin oxide product. There should be a whole number ratio between oxygen and tin. *(The hypothesis explains what is to be tested and will be written after reading the entire laboratory worksheet.)*

II. Equipment: *(Non-chemical equipment used in the experiment.)*
evaporating dish
watch glass
stirring rod
burner with ring stand, ring and wire gauze
forceps
beaker
balance

III. Reagents: *(A listing of chemicals used in the experiment with their amounts and any warnings.)*
tin metal (granulated) ~2 g.
5 cm³ (mL) nitric acid (HNO₃) **caution! severe burns**

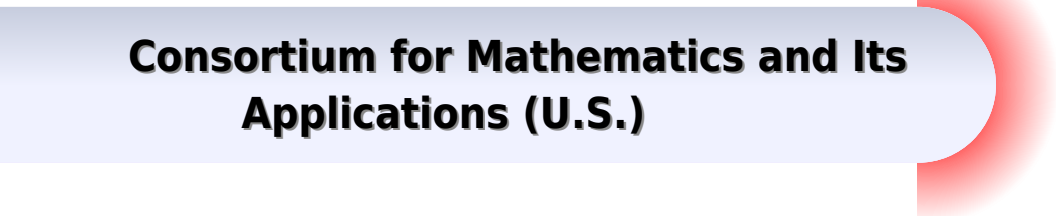
IV. Procedure: Each step of the procedure must be written here. You may paraphrase and shorten the explanations, but the reader must be able to perform the experiment from these instructions. *(The procedure must be read carefully before the lab begins. Drawings of the experimental setup may be included here. The teacher may make changes to the procedure; make sure that you write any changes down!)*

V. Data: *(If the laboratory report is handwritten use a ruler to draw data tables and graphs! Always include units with all data entries.)*

	Procedure	Trial 1	Trial 2
a	mass of dish, and watch glass	74.14 g	
b	mass of dish, glass and tin	76.20 g	
c	mass of tin =b-a	2.06 g	
d	moles of tin	.0173 mol	
e	mass of dish, glass, and product	76.76 g	
f	mass of oxygen =e-b	.56 g	
g	moles of oxygen	.0350 mol	
h	mole ratio	2.02 : 1	
i	accepted ratio	2 : 1	
j	% error	1.00 %	

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Table of Contents Sample Chemistry Laboratory Report

1. Understanding the eBook Sample Chemistry Laboratory Report
 - The Rise of Digital Reading Sample Chemistry Laboratory Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Sample Chemistry Laboratory Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sample Chemistry Laboratory Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sample Chemistry Laboratory Report
 - Personalized Recommendations
 - Sample Chemistry Laboratory Report User Reviews and Ratings
 - Sample Chemistry Laboratory Report and Bestseller Lists
5. Accessing Sample Chemistry Laboratory Report Free and Paid eBooks
 - Sample Chemistry Laboratory Report Public Domain eBooks
 - Sample Chemistry Laboratory Report eBook Subscription Services
 - Sample Chemistry Laboratory Report Budget-Friendly Options

6. Navigating Sample Chemistry Laboratory Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Sample Chemistry Laboratory Report Compatibility with Devices
 - Sample Chemistry Laboratory Report Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sample Chemistry Laboratory Report
 - Highlighting and Note-Taking Sample Chemistry Laboratory Report
 - Interactive Elements Sample Chemistry Laboratory Report
8. Staying Engaged with Sample Chemistry Laboratory Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sample Chemistry Laboratory Report
9. Balancing eBooks and Physical Books Sample Chemistry Laboratory Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sample Chemistry Laboratory Report
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Sample Chemistry Laboratory Report
 - Setting Reading Goals Sample Chemistry Laboratory Report
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Sample Chemistry Laboratory Report
 - Fact-Checking eBook Content of Sample Chemistry Laboratory Report
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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