

## Mathematics Formula Sheet & Explanation

The 2014 GED® Mathematical Reasoning test contains a formula sheet, which displays formulas relating to geometric measurement and certain algebra concepts. Formulas are provided to test-takers so that they may focus on *application*, rather than the *memorization*, of formulas.

### Area of a:

square	$A = s^2$
rectangle	$A = lw$
parallelogram	$A = bh$
triangle	$A = \frac{1}{2}bh$
trapezoid	$A = \frac{1}{2}h(b_1 + b_2)$
circle	$A = \pi r^2$

### Perimeter of a:

square	$P = 4s$
rectangle	$P = 2l + 2w$
triangle	$P = s_1 + s_2 + s_3$
Circumference of a circle	$C = 2\pi r$ OR $C = \pi d$ ; $\pi = 3.14$

### Surface area and volume of a:

rectangular prism	$SA = 2lw + 2lh + 2wh$	$V = lwh$
right prism	$SA = ph + 2b$	$V = bh$
cylinder	$SA = 2\pi rh + 2\pi r^2$	$V = \pi r^2 h$
pyramid	$SA = \frac{1}{2}ps + b$	$V = \frac{1}{3}bh$
cone	$SA = \pi rp + \pi r^2$	$V = \frac{1}{3}\pi r^2 h$
sphere	$SA = 4\pi r^2$	$V = \frac{4}{3}\pi r^3$

( $p$  = perimeter of base with area  $b$ ;  $\pi = 3.14$ )

### Data

mean	mean is equal to the total of the values of a data set, divided by the number of elements in the data set
median	median is the middle value in an odd number of ordered values of a data set, or the mean of the two middle values in an even number of ordered values in a data set

### Algebra

slope of a line	$m = \frac{y_2 - y_1}{x_2 - x_1}$
slope-intercept form of the equation of a line	$y = mx + b$
point-slope form of the equation of a line	$y - y_1 = m(x - x_1)$
standard form of a quadratic equation	$y = ax^2 + bx + c$
quadratic formula	$x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$
Pythagorean theorem	$a^2 + b^2 = c^2$
simple interest	$i = Prt$ ( $i$ = interest, $P$ = principal, $r$ = rate, $t$ = time)
distance formula	$d = rt$
total cost	total cost = (number of units) $\times$ (price per unit)

# Njask 2014 Formula Sheets

**Lei Huang**

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web introduction the physical and chemical behavior of compounds depends on the types of bonds that exist within the molecule and on the overall shape of the molecules the branch of chemistry that investigates how molecules are put together is

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web lewis structures represent the valence electrons present in the molecule as lone pairs unshared electrons drawn as dots or bonds shared electrons drawn as lines procedure for drawing lewis structures add up the total number of valence electrons that each atom contributes to the molecule polyatomic ion

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web the pcl 3 lewis structure has the typical case of phosphorus p in the center with 3 bonds to 3 other atoms phosphorus is from the same column as nitrogen in the periodic table meaning that p and n generally have the same bonding structure note the lone pair dots without bonds on top of p just like for n in the previous example for nh 3 chlorine cl

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web step 1 figure out how many electrons the molecule must have based on the number of valence electrons in each atom when drawing the structure of an ion be sure to add subtract electrons to account for the charge step 2 connect the atoms to each other with single bonds to form a skeleton structure

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web lewis structures a lewis structure is a representation of covalent molecules or polyatomic ions where all the valence electrons are shown distributed about the bonded atoms as either shared electron pairs bond pairs or unshared electron

pairs lone pairs a shared pair of electrons is represented as a short line a single bond

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2 total number of valence electrons 24 cad engineered 3d sketch model show dipole arrows lewis structure is there a polar  
bond in this molecule yes or no vsepr shape name trigonal planar bond angles 120

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web jun 8 2022 lewis structure in chemistry lewis structures also called electron dot structures or electron dot diagrams are  
diagrams that show the bonding between atoms of a molecule and the lone pairs of

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web this exercise required us to take the formulas in the data table and calculate the number of valence electrons using the  
periodic table draw its lewis structure on paper build the molecular model and record a photo list the number of bonds and  
lone pairs surrounding the central atom and lastly identify the structure geometry