

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete ionic equation.

Vocabulary

- 1. chemical equation
- 2. balanced equation
- 3. catalyst
- 4. coefficient
- 5. skeletal equation

Part A Completion

Write the coefficient for each in these chemical equations, and balance and label every chemical reaction. Indicate the physical states of the reactants and products.

1. A chemical reaction is usually represented by a chemical _____ 2. _____
 The chemical formula for a chemical compound is _____ 3. _____ The mass
 number for an element is represented by the _____ 4. _____ An element's atomic
 mass is represented by _____ 5. _____ A chemical equation is written with
 reactants on the left and products on the right _____ 6. _____ A chemical equation is
 balanced when the number of atoms of each element is the same on both sides of the
 equation. The equation must be balanced by changing the _____ 7. _____ or
 the chemical formula of the reactants.

Chemical reactions are written with physical states indicated by a
 symbol. The symbol for liquid is _____ 8. _____ for a solid _____ 9. _____ for a gas
 _____ 10. _____ A chemical equation is balanced _____ 11. _____ if _____ 12. _____ or
 mass is conserved. A chemical equation is balanced when the number of atoms of each
 element _____ 13. _____

- 1. equation
- 2. formula
- 3. atomic mass
- 4. mass
- 5. coefficient
- 6. balanced
- 7. coefficients
- 8. l
- 9. s
- 10. g
- 11. balanced
- 12. mass
- 13. conserved

Section Review 11 1 Describing Chemical Reactions

Sasha Hu



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Applied Mechanics Reviews, 1973 **Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *ERDA Energy Research Abstracts* United States. Energy Research and Development Administration, 1977 **ERDA Energy Research Abstracts** United States. Energy Research and Development Administration. Technical Information Center, 1977 *Voet's Principles of Biochemistry* Donald Voet, Charlotte W. Pratt, Judith G. Voet, 2018 Voets Principles of Biochemistry Global Edition addresses the enormous advances in biochemistry particularly in the areas of structural biology and bioinformatics It provides a solid biochemical foundation that is rooted in chemistry to prepare students for the scientific challenges of the future New information related to advances in biochemistry and experimental approaches for studying complex systems are introduced Notes on a variety of human diseases and pharmacological effectors have been expanded to reflect recent research findings While continuing in its tradition of presenting complete and balanced coverage this Global Edition includes new pedagogy and enhanced visuals that provide a clear pathway for student learning 4e de couverture Energetic Principles of Chemical Reactions Jack Simons, 1983 **Energy Research Abstracts**, 1990 **Weather and Climate** David Money, 2000

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