Name of Acid	Acid		Base	Ka
Hydriodic	HI		H+ + I-	>> 1
Hydrobromic	HBr	>	H+ + Br-	>> 1
Hydrochloric	HCI	>	H+ + CI-	>> 1
Sulfuric	H ₂ SO ₄	>	H+ + HSO4-	> 1
Nitric	HNO ₃		H+ + NO3-	> 1
Hydronium Ion	H ₃ O+	-	H+ + H ₂ O	5.50 X 10 ¹
Oxalic	H2C2O4		H+ + HC2O4-	5.90 X 10 ⁻²
Sulfurous	H ₂ SO ₃	-	H+ + HSO3-	1.41 X 10 ⁻²
Hydrogen Sulfate Ion	HSO ₄	-	H+ + SO ₄ 2-	1.20 X 10 ⁻²
Phosphoric	H ₃ PO ₄	-	H+ + H2PO4-	7.52 X 10 ⁻³
Nitrous	HNO ₂	-	H+ + NO2-	5.62 X 10 ⁻⁴
Hydrofluoric	HF		H+ + F-	3.53 X 10 ⁻⁴
Formic	HCO ₂ H	-	H+ + HCO2-	1.78 X 10 ⁻⁴
Benzoic	C ₆ H ₅ CO ₂ H	~	H+ + C6H5CO2-	6.46 X 10 ⁻⁵
Acetic	CH ₃ CO ₂ H		H+ + CH3CO2-	1.76 X 10 ⁻⁵
Carbonic	H ₂ CO ₃	-	H+ + HCO3-	4.30 X 10 ⁻⁷
Hydrogen Sulfite Ion	HSO ₃	-	H+ + SO ₃ 2-	1.02 X 10 ⁻⁷
Dihydrogen Phosphate Ion	H ₂ PO ₄	-	H+ + HPO42-	6.23 X 10 ⁻⁸
Boric	H ₃ BO ₃	-	H+ + H2BO3-	5.79 X 10 ⁻¹⁰
Ammonium Ion	NH ₄ +	~	H* + NH ₃	5.64 X 10 ⁻¹⁰
Hydrogen Carbonate Ion	HCO ₃ -	-	H+ + CO32-	5.61 X 10 ⁻¹¹
Monohydrogen Phosphate Ion	HPO ₄ 2-	-	H+ + PO43-	2.20 X 10 ⁻¹³
Water	H ₂ O		H+ + OH-	1.00 X 10 ⁻¹⁴

All reactions assume water as a reactant. $HA + H_2O \Longrightarrow H_3O^+ + A^-$ Innovating Science®

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 $K_a = [H^+][A^-] / [HA]$ $pK_a = - \log K_a$ $pH = - \log [H^+]$

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