

## REVIEW: Chapter 8 - Covalent Bonding

- Typically, how many bonds & lone pairs will each have when covalently bonding?
  - Boron – 3 bonds; 0 lone pairs
  - Silicone – 4 bonds; 0 lone pairs
  - Neon – 0 bonds; 4 lone pairs
  - Iodine – 1 bond; 3 lone pairs
  - Nitrogen – 3 bonds; 1 lone pair
  - Oxygen – 2 bonds; 2 lone pairs
- Which elements do not follow the octet rule? (Hint – there are four we went over)
  - H & B never follow the octet rule.
  - P & S sometimes don't follow the octet rule.
- What type of bond will each combination of atoms form?
  - F & N – 2 nonmetals, therefore covalent
  - Al & O – 1 metal/1 nonm., therefore ionic
  - Na & S – 1 metal/1 nonm., therefore ionic
  - B & Cl – 2 nonmetals, therefore covalent
  - Ca & I – 1 metal/1 nonm., therefore ionic
  - Fe & Cu – 2 metals, therefore metallic
- What is the difference between:
  - Molecules & Compounds?  
Compounds can be made using all types of bonds (ionic, covalent or metallic), but molecules can only be made using covalent bonds
  - Monatomic elements & Diatomic molecules?  
Monatomic elements exist as single atoms; diatomic molecules are made up of two of the same atom
  - Structural formulas, Empirical formulas, & Molecular formulas?  
All three are for covalent bonds only! Molecular show the actual number of atoms in the molecule; Empirical give the lowest whole number ratio; Structural formulas are not really formulas, they are drawing of the molecule.
  - Polar bonds & Non-polar bonds?  
Both describe covalent bonds! In a polar bond the electrons are NOT shared equally; in a non-polar bond electrons are shared equally
- Classify each of the following as an atom or a molecule:
  - Be, atom
  - CO, molec.
  - N<sub>2</sub>, molec.
  - H<sub>2</sub>O, molec.
  - Ne, atom

# Section 8 Covalent Bonding Workbook Answers

**Tao Wei**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name Tao Wei.

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