

1. In the list below, circle the properties of acids and box the properties of bases.

Sour taste	Feel slippery	Contains hydroxide ions
Turns litmus blue	Reacts with metals to form H_2	Turns litmus red
Bitter taste	Contains hydrogen ions	Corrosive

2. Complete the chart below:

Acid-Base Definitions		
Type	Acid Definition	Base Definition
Arrhenius	Donates an H^+ ion in solution	Donates an OH^- ion in solution
Bronsted-Lowry	Donates an H^+ ion	Accepts an H^+ ion
Lewis	Accepts an electron pair	Donates an electron pair

3. What term(s) describe a substance that can act as an acid or a base? **Amphoteric or amphiprotic**
4. Neutralization reactions always produce a **SALT** and **WATER**
5. If the $[H^+]$ is greater than the $[OH^-]$, the solution is acidic basic neutral
6. If the $[OH^-]$ is greater than the $[H^+]$, the solution is acidic basic neutral
7. If the $[H^+]$ is equal to the $[OH^-]$, the solution is acidic basic neutral
8. If the pH is 7, the solution is acidic basic neutral
9. If the pH is greater than 7, the solution is acidic basic neutral
10. If the pH is less than 7, the solution is acidic basic neutral
11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?



$$(19.2 \text{ g } HClO_3) \times (1 \text{ mol } HClO_3/84 \text{ g } HClO_3) \times (1 \text{ mol } LiOH/1 \text{ mol } HClO_3) = \underline{0.228 \text{ mol } LiOH}$$

12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?



$$(132.4 \text{ g } HNO_3) \times (1 \text{ mol } HNO_3/63 \text{ g } HNO_3) \times (1 \text{ mol } Ca(OH)_2/2 \text{ mol } HNO_3) \times (74 \text{ g } Ca(OH)_2/1 \text{ mol } Ca(OH)_2) = \underline{77.76 \text{ g } Ca(OH)_2}$$

Section 8 3 Properties Of Acids And Base Answers

JR Anderson



Section 8 3 Properties Of Acids And Base Answers:

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Solomons, Craig B. Fryhle, Scott A. Snyder, 2022-01-19 Organic Chemistry 13th Edition continues Solomons Fryle and Snyder's tradition of excellence in teaching and preparing students for success in both the classroom and beyond Central to the authors is their approach in emphasizing organic chemistry's relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us Science in Your World: Teacher edition Jay K. Hackett, 1991

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