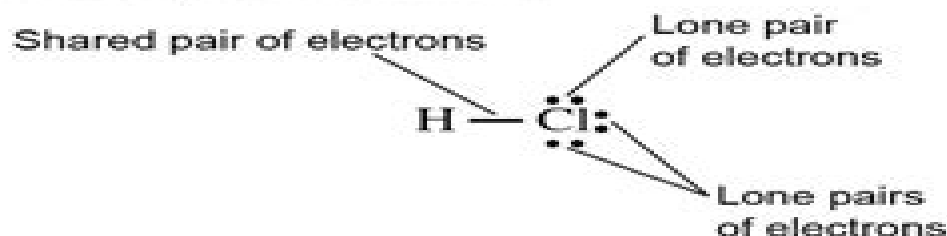


Ch. 6 - Covalent Bonding

Section 6.1 - Formation of Covalent Bonds

- Electrons being attracted to two nuclei at the same time - covalent bonding
 - Then electrons are shared between the two nuclei making the molecules more stable
- If bond length is too far apart - the attractive forces are too small and results in no bond
- If bond lengths are too close - the repulsive forces are too large and push the atoms apart and results in no bond
- Optimal separation is when energy is at its lowest
- Every bond has a predicted length adhered from the atomic radii
- Want to be the most stable (completely made across the periodic table achieving noble gas status, $ns^2 np^6$)
 - octet rule - achieving 8 valence electrons
- Only electrons that can be shared is valence electron
 - - bold dash represents covalent bonds



Section 6.2 - Lewis Structures

- Number of electrons in lewis structure shows the total number of valence electrons in the atoms of a molecule
 - Octet (noble gas - 8 valence electrons)
 - Duet (add two electrons)

Section 6 Covalent Bonding Lincoln

Lei Huang



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