I. ONE REACTANT IN EXCESS - COMBUSTION

You burn an excess of butane (C₄H₁₀) in a limited amount of oxygen (O₂). If 67.2 L of CO₂ gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:

STEP 2: Use a molar conversion to convert quantity of product given to moles:

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol CO}_2$$
 * (13 mol O_2) = 4.875 mol O_2
 (8 mol CO_2)

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol } O_2 \ (22.4 \text{ L} \cdot O_2) = 109.2 \text{ L} \cdot O_2 = 109 \text{ L} \cdot O_2$$

(1 mol O_2)

II. ONE REACTANT IN EXCESS - SINGLE DISPLACEMENT

14 g of lithium is combined with excess K₂SO₄. What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP I: Balance the reaction:

Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

2 mol Li *
$$(1 \text{ mol Li}_2SO_4) = 1 \text{ mol Li}_2SO_4$$

(2 mol Li)

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

Section 2 Stoichiometry Answers

Albert A Gayle

Section 2 Stoichiometry Answ	vers:
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This book delves into Section 2 Stoichiometry Answers. Section 2 Stoichiometry Answers is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Section 2 Stoichiometry Answers, encompassing both the fundamentals and more intricate discussions.

- 1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Section 2 Stoichiometry Answers
 - Chapter 2: Essential Elements of Section 2 Stoichiometry Answers
 - Chapter 3: Section 2 Stoichiometry Answers in Everyday Life
 - Chapter 4: Section 2 Stoichiometry Answers in Specific Contexts
 - ∘ Chapter 5: Conclusion
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- 3. In chapter 2, the author will delve into the foundational concepts of Section 2 Stoichiometry Answers. The second chapter will elucidate the essential principles that must be understood to grasp Section 2 Stoichiometry Answers in its entirety.
- 4. In chapter 3, the author will examine the practical applications of Section 2 Stoichiometry Answers in daily life. This chapter will showcase real-world examples of how Section 2 Stoichiometry Answers can be effectively utilized in everyday scenarios.
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- 6. In chapter 5, the author will draw a conclusion about Section 2 Stoichiometry Answers. This chapter will summarize the key points that have been discussed throughout the book.
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