

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                       |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base     |
| 4. autoionization | 5. acid rain                | 6. acid rain solution |
| 7. autoionization | 8. acid rain                | 9. pH                 |

## Key Equation

- |                                    |                                    |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

## Part A Completion

Read the paragraph and fill in the blanks by understanding of the concepts presented. Then check answers in Part C.

When pure water is \_\_\_\_\_, it ionizes. Hydrogen ion concentration is \_\_\_\_\_ the concentration of hydroxide ion concentration in pure water. \_\_\_\_\_, the concentration of hydroxide ion is \_\_\_\_\_ the concentration of hydrogen ion. \_\_\_\_\_, the concentration of hydroxide ion is \_\_\_\_\_ the concentration of hydrogen ion. \_\_\_\_\_, the concentration of hydroxide ion is \_\_\_\_\_ the concentration of hydrogen ion. \_\_\_\_\_, the concentration of hydroxide ion is \_\_\_\_\_ the concentration of hydrogen ion.

\_\_\_\_\_ is the concentration of hydrogen ion. \_\_\_\_\_ is the concentration of hydroxide ion. \_\_\_\_\_ is the concentration of hydrogen ion. \_\_\_\_\_ is the concentration of hydroxide ion. \_\_\_\_\_ is the concentration of hydrogen ion. \_\_\_\_\_ is the concentration of hydroxide ion.

1. autoionization
2.  $1 \times 10^{-7} \text{ M}$
3.  $1 \times 10^{-7} \text{ M}$
4.  $1 \times 10^{-7} \text{ M}$
5.  $1 \times 10^{-7} \text{ M}$
6.  $1 \times 10^{-7} \text{ M}$
7.  $1 \times 10^{-7} \text{ M}$
8.  $1 \times 10^{-7} \text{ M}$
9.  $1 \times 10^{-7} \text{ M}$
10.  $1 \times 10^{-7} \text{ M}$
11.  $1 \times 10^{-7} \text{ M}$
12.  $1 \times 10^{-7} \text{ M}$
13.  $1 \times 10^{-7} \text{ M}$
14.  $1 \times 10^{-7} \text{ M}$
15.  $1 \times 10^{-7} \text{ M}$

## Section 19 2 Hydrogen And Ions Acidity Answer Key

**Puja Mehta**



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