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# Acids and Bases Reinforcement Activity

Write "A" on the line if the following substance is an acid. Write "B" on the line if it is a base. If the substance is generally not an acid or a base, write "N" for neither. Most bases are metal hydroxides (OFI), metal oxides (OF), or armonium (NH<sub>4</sub>") hydroxides/oxides.

L H <sub>2</sub> O	15. Ca(OH) <sub>k</sub>	29. HCi
z. H <sub>2</sub> SO <sub>4</sub>		30. Li <sub>2</sub> O
s. NivOH	17. HNOs	31. HBr
4. MgO	18. CeHoOs	32. NapO
s. HCN	19. HNOs	33. HI
S. HCIO4	20. HCx0x0x	34. HF
7. HCKO <sub>3.</sub>	21. PO <sub>4</sub> *	35. K <sub>2</sub> O
k HCIO <sub>1</sub>	22 Mg <sup>3*</sup>	36. H <sub>3</sub> S
9. SOV2	23. H <sub>2</sub> C <sub>2</sub> O <sub>4</sub>	37. C <sub>2</sub> H <sub>4</sub>
IO. HCIO	24 H <sub>2</sub> BO <sub>3</sub>	38. C <sub>1</sub> H <sub>4</sub>
II. NieOH	25. HiSOs	39. KMeOv
12. LiOH	26. SrO	40. HxO0v
13. Bu(OH)s		41. HsPOs
14. NHAOH	28 Rb <sub>1</sub> O	42. SerOH)s

Complete and balance the acid-base neutralization reactions. Write "s" under the acid (reactant), and write "b" under the base (reactant). Write "s" under the salt, which is the new ionic compound produced. Hint: carbonate ions (CO<sub>3</sub><sup>3</sup>) will split spart to produce CO<sub>2</sub> gas and to help from water.

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