

# DESCRIBING CHEMICAL REACTIONS

## Section Review - Objectives

- Explain how to write a word equation
- Describe how to write a skeleton equation
- List the steps for writing a complete chemical equation

## Vocabulary

- chemical equation
- skeleton equation
- catalyst
- coefficients
- balanced equation

## Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

A chemical reaction can be concisely represented by a chemical \_\_\_\_1\_\_\_\_. The substances that undergo a chemical change are the \_\_\_\_2\_\_\_\_. The new substances formed in a chemical reaction are the \_\_\_\_3\_\_\_\_. In accordance with the law of conservation of \_\_\_\_4\_\_\_\_, a chemical equation must be balanced. When balancing an equation, you place \_\_\_\_5\_\_\_\_ in front of reactants and products so that the same number of atoms of each \_\_\_\_6\_\_\_\_ are on each side of the equation. An equation must never be balanced by changing the \_\_\_\_7\_\_\_\_ in the chemical formula of a substance.

Special symbols are used to show the physical state of a substance in a reaction. The symbol for a liquid is \_\_\_\_8\_\_\_\_; for a solid, \_\_\_\_9\_\_\_\_; and for a gas, \_\_\_\_10\_\_\_\_. A substance dissolved in water is designated \_\_\_\_11\_\_\_\_. If a \_\_\_\_12\_\_\_\_ is used to increase the rate of a chemical reaction, its formula is written above the arrow.

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## Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- \_\_\_\_\_ 13. In an equation, a substance is shown to be in the gaseous state by placing an upward-pointing arrow after its formula.
- \_\_\_\_\_ 14. The symbol  $\Delta$  placed over the arrow in an equation means that heat is supplied to the reaction.
- \_\_\_\_\_ 15. Atoms are destroyed in a chemical reaction.
- \_\_\_\_\_ 16. A skeleton equation is not a balanced equation.

## Part D Questions and Problems

24. Write a balanced equation for each of these chemical reactions.
  - a. Aluminum reacts with aqueous hydrochloric acid to form hydrogen gas and aqueous aluminum chloride.
  - b. Acetylene gas ( $C_2H_2$ ) burns in a welding torch with oxygen to form carbon dioxide gas and water vapor.

# Section 1 2 Describing Chemical Reactions Answer Sheet

**Christian G. Meyer**



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## **Table of Contents Section 1 2 Describing Chemical Reactions Answer Sheet**

1. Understanding the eBook Section 1 2 Describing Chemical Reactions Answer Sheet
  - The Rise of Digital Reading Section 1 2 Describing Chemical Reactions Answer Sheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 1 2 Describing Chemical Reactions Answer Sheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 1 2 Describing Chemical Reactions Answer Sheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 1 2 Describing Chemical Reactions Answer Sheet
  - Personalized Recommendations
  - Section 1 2 Describing Chemical Reactions Answer Sheet User Reviews and Ratings

- Section 1 2 Describing Chemical Reactions Answer Sheet and Bestseller Lists
- 5. Accessing Section 1 2 Describing Chemical Reactions Answer Sheet Free and Paid eBooks
  - Section 1 2 Describing Chemical Reactions Answer Sheet Public Domain eBooks
  - Section 1 2 Describing Chemical Reactions Answer Sheet eBook Subscription Services
  - Section 1 2 Describing Chemical Reactions Answer Sheet Budget-Friendly Options
- 6. Navigating Section 1 2 Describing Chemical Reactions Answer Sheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Section 1 2 Describing Chemical Reactions Answer Sheet Compatibility with Devices
  - Section 1 2 Describing Chemical Reactions Answer Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Section 1 2 Describing Chemical Reactions Answer Sheet
  - Highlighting and Note-Taking Section 1 2 Describing Chemical Reactions Answer Sheet
  - Interactive Elements Section 1 2 Describing Chemical Reactions Answer Sheet
- 8. Staying Engaged with Section 1 2 Describing Chemical Reactions Answer Sheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Section 1 2 Describing Chemical Reactions Answer Sheet
- 9. Balancing eBooks and Physical Books Section 1 2 Describing Chemical Reactions Answer Sheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Section 1 2 Describing Chemical Reactions Answer Sheet
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Section 1 2 Describing Chemical Reactions Answer Sheet
  - Setting Reading Goals Section 1 2 Describing Chemical Reactions Answer Sheet
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 1 2 Describing Chemical Reactions Answer Sheet
  - Fact-Checking eBook Content of Section 1 2 Describing Chemical Reactions Answer Sheet
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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