

6

THE PERIODIC TABLE

Chapter 6 Review Package - Part 1

A. Matching

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

Column A

- _____ 1. metal
- _____ 2. halogen
- _____ 3. inner transition metal
- _____ 4. representative element
- _____ 5. ~~ionization energy~~
- _____ 6. alkaline earth metal
- _____ 7. cation
- _____ 8. electronegativity
- _____ 9. transition metal
- _____ 10. noble gas

Column B

- a. positively charged ion
- b. an element whose highest occupied *s* or *p* sublevels are partially filled
- c. ~~the energy needed to remove an electron from an atom in the gaseous state~~
- d. good conductor of heat and electric current
- e. an element whose highest occupied *s* sublevel and a nearby *d* sublevel contain electrons
- f. any nonmetal in Group 7A
- g. ~~the tendency of an atom to attract electrons when the atom is in a compound~~
- h. an element in which the highest occupied *s* and *p* sublevels are filled
- i. an element in Group 2A
- j. an element whose highest occupied *s* sublevel and a nearby *f* sublevel contain electrons

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. In the periodic table, there is a periodic pattern in the physical and chemical properties of elements when they are arranged in order of
 - a. increasing atomic mass.
 - b. increasing electronegativity.
 - c. increasing atomic radius.
 - d. increasing atomic number.

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Karin Nielsen-Saines

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