

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete ionic equation.

Vocabulary

- 1. chemical equation
- 2. balanced equation
- 3. catalyst
- 4. coefficient
- 5. skeletal equation

Part A Completion

Write the coefficient for each in the balanced equation, and balance the equation by changing the coefficient for each substance. Write balanced ionic equations for each.

1. A chemical equation is always represented by a balanced _____ 2. _____
 The chemical formula for a chemical compound is _____ 3. _____
 Chemical formulas for compounds are the _____ 4. _____
 The law of conservation of _____ 5. _____
 A chemical equation is balanced when the number of _____ 6. _____
 atoms of each element is the same on both sides of the equation. _____ 7. _____
 To balance a chemical equation, you place _____ 8. _____
 coefficients in front of the chemical formulas. _____ 9. _____
 The _____ 10. _____
 _____ 11. _____
 _____ 12. _____
 _____ 13. _____
 _____ 14. _____
 _____ 15. _____
 _____ 16. _____
 _____ 17. _____
 _____ 18. _____
 _____ 19. _____
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_____ 21. _____
 _____ 22. _____
 _____ 23. _____
 _____ 24. _____
 _____ 25. _____
 _____ 26. _____
 _____ 27. _____
 _____ 28. _____
 _____ 29. _____
 _____ 30. _____

- 1. equation
- 2. formula
- 3. chemical
- 4. law
- 5. mass
- 6. balanced
- 7. coefficient
- 8. atoms
- 9. balanced
- 10. balanced
- 11. balanced
- 12. balanced
- 13. balanced
- 14. balanced
- 15. balanced
- 16. balanced
- 17. balanced
- 18. balanced
- 19. balanced
- 20. balanced

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