

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev's original short periodic table gave better chemical properties of elements appeared in regular intervals and whereas elements were arranged according to increasing
 - (a) density
 - (b) atomic number
 - (c) atomic mass
 - (d) electronegativity
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a column
 - (c) adding a new group
 - (d) adding a new period
5. _____ The noble gases that property of the noble gases is that they are
 - (a) unstable
 - (b) inert
 - (c) radioactive
 - (d) highly reactive
6. _____ Likewise, the first column in Group 1, has characteristic number of 1. The second column in the group has characteristic number of
 - (a) 2
 - (b) 10
 - (c) 18
 - (d) 32
7. How groups in the periodic table are numbered? (a) and (b) atomic number of (b)
 - (a) 1 to 18
 - (b) 1 to 10
 - (c) 1 to 18
 - (d) 1 to 10
8. How many periods are there in the periodic table?
 - (a) 18
 - (b) 10
 - (c) 18
 - (d) 10
9. How many elements are in the periodic table?
 - (a) 118
 - (b) 110
 - (c) 118
 - (d) 110
10. What is the periodic table's main purpose?
 - (a) to show the relationship between elements and their properties
 - (b) to show the relationship between elements and their properties
 - (c) to show the relationship between elements and their properties
 - (d) to show the relationship between elements and their properties

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PT Brinkman



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