

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|---|------------------|
| 1. autoionization | 2. hydronium ion (H_3O^+) | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. acid solution |
| | | 7. pH |

Key Equation

- | | |
|------------------------------------|--|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |

Part A Completion

Write the equation for the autoionization of water. What is the concentration of H^+ in pure water? What is the concentration of OH^- in pure water?

What is the pH of a solution with $[\text{H}^+] = 1.0 \times 10^{-7} \text{ M}$?

What is the pH of a solution with $[\text{OH}^-] = 1.0 \times 10^{-7} \text{ M}$?

What is the pH of a solution with $[\text{H}^+] = 1.0 \times 10^{-7} \text{ M}$?

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What is the pH of a solution with $[\text{OH}^-] = 1.0 \times 10^{-7} \text{ M}$?

What is the pH of a solution with $[\text{H}^+] = 1.0 \times 10^{-7} \text{ M}$?

What is the pH of a solution with $[\text{OH}^-] = 1.0 \times 10^{-7} \text{ M}$?

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

2. $1.0 \times 10^{-7} \text{ M}$

3. $1.0 \times 10^{-7} \text{ M}$

4. $1.0 \times 10^{-7} \text{ M}$

5. $1.0 \times 10^{-7} \text{ M}$

6. $1.0 \times 10^{-7} \text{ M}$

7. $1.0 \times 10^{-7} \text{ M}$

8. $1.0 \times 10^{-7} \text{ M}$

9. $1.0 \times 10^{-7} \text{ M}$

10. $1.0 \times 10^{-7} \text{ M}$

11. $1.0 \times 10^{-7} \text{ M}$

12. $1.0 \times 10^{-7} \text{ M}$

Section 19 2 Hydrogen Ions And Acidity Answer Key

**Linda Lane Lilley, Julie S. Snyder, Shelly
Rainforth Collins**



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biophysical methods and their applications **Cumulated Index Medicus**, 1980 **Study Guide for Lewis'**

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