RATES OF REACTION

Section A: Summary Notes

Energy Changes during Chemical Reactions

- Most reactions do not begin until an amount of energy (activation energy) has been added to the reaction mixture.
- The activation energy is often called the 'energy hill' which must be 'overcome' by the addition
 of this amount of energy before a reaction can take place.
- When activation energy is added to the reactants, a so-called activated complex is formed.
- The activated complex is temporary, unstable, high-energy composition of atoms, which
 represents a transition state between reactants and the products.
- When the activated complex is formed during a reaction, this complex can lead either to the formation of new bonds, i.e molecules of the products, or to re-formation of the old bonds, thereby returning to being reactants.
- For the reaction:

AB • C → A • BC

The formation of an activated complex as a transitional step can be represented as follows:

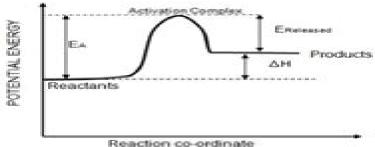
AB • C → [ABC] → A • BC

 When an activated complex is formed during a reaction, this complex can lead either to the formation of new bonds, i.e. molecules of the products or to the re-formation of the old bonds, thereby returning to being the reactions. This is reversibility for the reaction.

Heat of Reaction (Enthalpy)

Heat of the reaction (ΔH) is the difference between the energy of the products and the energy
of the reactions.

For an endothermic reaction, Engages Engages, therefore AH is positive.



Presidence of Co-order Hard

Note: In a reversible reaction, the energy released in forming the products in the forward reaction is the same as the activation energy (Ex) of the reverse reaction.

For an exothermic reaction, Epoplets Emergen, AH is negative.

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