



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following as ionic or molecular.

- a. NaCl (ionic) b. CH_4 (molecular) c. H_2O (molecular)
 d. KNO_3 (ionic) e. H_2SO_4 (molecular)

2. Which of the following are molecular compounds?

- a. H_2O b. CO_2 c. H_2O_2
 d. NaCl e. H_2SO_4 f. H_2O_2

3. Which of the following are molecular compounds? Write the molecular formula.

a. Carbon dioxide is a CO_2 molecular compound. The molecular formula is CO_2 .
 b. Sulfur dioxide is a SO_2 molecular compound. The molecular formula is SO_2 .
 c. Hydrogen peroxide is a H_2O_2 molecular compound. The molecular formula is H_2O_2 .

d. Water is a H_2O molecular compound. The molecular formula is H_2O .
 e. Hydrogen sulfide is a H_2S molecular compound. The molecular formula is H_2S .

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .



4. Draw the Lewis structure for acetylene, C_2H_2 .

5. Draw the Lewis structure for hydrogen cyanide, HCN . Is this a molecular compound or ionic?



6. Draw the Lewis structure for hydrogen cyanide, HCN .

SECTION 8.3 BONDING: THE ORBITAL

1. Which orbital has two lobes, one above and one below the nucleus? Is it a bonding orbital or an antibonding orbital?

2. Which orbital has two lobes, one to the left and one to the right of the nucleus? Is it a bonding orbital or an antibonding orbital?

Section 8 1 Covalent Bonding Pearson Education Key

Daniela Niemeyer



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