

Chapter 2 Minerals

Section 2.1 Matter

pages 34–47

This section discusses the relationship between minerals and elements. It explains the parts of an atom and defines ions, isotopes, compounds, and chemical bonds.

Reading Strategy (page 34)

Comparing and Contrasting As you read, complete the organizer to compare and contrast protons, neutrons, and electrons. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Protons	Electrons	Neutrons
	Differences	Neutron is not charged.
Similarities		

Elements and the Periodic Table (pages 34–37)

1. ☐ A substance that cannot be broken down into simpler substances is called a(n) _____. Circle the correct answer.
substance compound element
2. The document in which elements are organized by their properties is known as the _____. Circle the correct answer.
periodic table common elements Earth materials

Atoms (pages 38, 39)

3. ☐ What is an atom? _____
4. The atomic number of boron is 5. What does this tell you about an atom of boron? _____

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