



Unit 1: Chemical Changes

Chemical Changes

Answer the following questions using the following equation:



1. What are the products of this chemical reaction?

The products of this chemical reaction are the formation of water and oxygen. In this chemical reaction, the reactants are hydrogen peroxide (H_2O_2) and the products are water (H_2O) and oxygen (O_2). The reaction is exothermic, meaning it releases heat and energy.

2. What is the reactant in this chemical reaction? Is it a solid, liquid, gas, or a mixture? Is it a pure substance or a mixture?

3. What is the physical state of the reactant and the products?

The reactant, hydrogen peroxide, is a liquid. The products, water and oxygen, are a liquid and a gas, respectively. The reaction is exothermic, meaning it releases heat and energy. The reaction is also a decomposition reaction, meaning it breaks down a single substance into two or more simpler substances.

4. How many moles of reactant are there? How many moles of product are there? How many moles of gas are there? How many moles of liquid are there?

2 moles



5. What is the chemical formula of the reactant? What is the chemical formula of the product?

6. What is the chemical formula of the reactant? What is the chemical formula of the product?

7. What are the physical states of the reactant and the products?

Reactant: liquid
Products: liquid and gas

8. What is the physical state of the reactant and the product?

9. What is the chemical formula of the reactant and the product?

Reactant: H_2O_2 (liquid)
Products: H_2O (liquid) and O_2 (gas)

10. How many moles of reactant are there? How many moles of product are there? How many moles of gas are there? How many moles of liquid are there?

Section 1 Reinforcement Chemical Changes

R Barnett



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