

## 5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

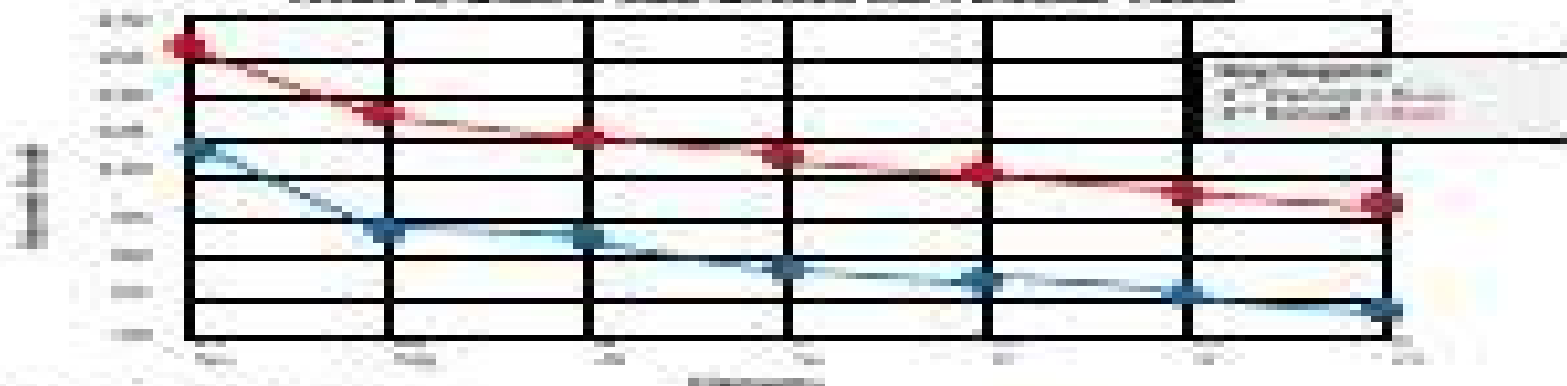
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property. Sketch it's data points graphically.
2. Is a slight downward trend in this graph Period? If not, fill in the other graphs.
3. Labelled as down or up include metallic
4. Answer all questions.

	1	2	3	4	5	6	7	8
1	1H	2He						
2	3Li	4Be	5B	6C	7N	8O	9F	10Ne
3	11Na	12Mg	13Al	14Si	15P	16S	17Cl	18Ar
4	19K	20Ca	21Sc	22Ti	23V	24Cr	25Mn	26Fe
5	27Co	28Ni	29Cu	30Zn	31Ga	32Ge	33As	34Se
6	35Br	36Kr	37Rb	38Sr	39Y	40Zr	41Nb	42Mo
7	43Tc	44Ru	45Rh	46Pd	47Ag	48Cd	49In	50Sn

Now label the properties above in the table below:

Trend in Atomic Size across the Periodic Table



### Electronegativity Questions

1. Following the trend in atomic size across the periodic table, decrease
2. This means two atoms that are most different properties, differences they are all different (making electronegativity trend in electronegativity higher)

### Electronegativity Questions

1. Electronegativity trend in the periodic table is decrease (e.g. Na, Cl, Br, I, At). The trend in electronegativity across the periodic table is decrease
2. This means two atoms that are most different properties, differences they are all different (making electronegativity trend in electronegativity higher)

# Periodicity Activity Okstate Answer Key

**A Loxley**



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