

Directions: Choose the letter that corresponds to the correct answer.

1. Which of the following can be considered as Dobereiner's triad?
 - a. He: 4.0, Li: 6.9, Na: 23.0
 - b. N: 14.0, P: 40.0, At: 217.0
 - c. Ne: 20.2, Ar: 39.9, Kr: 83.8
 - d. Ca: 40.1, Sr: 87.6, Ba: 137.3
2. Which of the following elements does not belong to the group?
 - a. N
 - b. S
 - c. Po
 - d. Te
3. Which of the following is/are TRUE about boron and oxygen?
 - a. both are metalloids
 - b. boron is smaller than oxygen
 - c. boron is more electronegative than oxygen
 - d. higher energy is required to eject an electron from the valence shell of boron
 - e. all of the above
 - f. none of the above

For 4 and 5, consider Na, Mg, Al, and Si.

4. The element that ejects one electron from its valence shell most readily is _____.
 - a. Na
 - b. Mg
 - c. Al
 - d. Si
5. Which element has the most negative electron affinity?
 - e. Na
 - f. Mg

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Princeton Review (Firm)



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