

Observation Table :

Mass of clean, empty crucible and lid (g)	31.072
Mass of crucible, lid, and magnesium (g)	31.638
Mass of crucible, lid, and magnesium oxide (g)	31.992
Mass of magnesium oxide produced (g)	0.906 g
Mass of oxygen that must have reacted (g)	0.336 g
Percent mass of magnesium in magnesium oxide $\frac{m_{\text{(Mg)}}}{m_{\text{(MgO)}}} \times 100\%$	62.9 % $\approx 63\%$

$$\frac{0.570}{0.906} \times 100 = 62.91$$

Analyze and Interpret

1. What mass of magnesium was used in the reaction?
2. What mass of magnesium oxide was produced?
3. Calculate the mass of oxygen that reacted with the magnesium.

Conclude and Communicate

4. Use the data to calculate the percentage composition of magnesium oxide.
5. Calculate the percentage error (PE) using the difference between the experimental mass percent (EP) of magnesium and the actual mass percent (AP) of magnesium.
6. Suppose that some magnesium oxide smoke escaped during the investigation. Would the Mg : O ratio have increased, decreased, or remained unchanged? Explain your answer using sample calculations.
7. How would the value you calculated for the percentage composition of magnesium oxide have been affected if all the magnesium in the crucible had not reacted? Support your answer with sample calculations.

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