#### **Chapter 6: Chemical Bonding Guided Notes**

#### B. Why do Atoms Combine?

- A. Chemical Formula
  - The chemical formula shows:
    - a. The \_\_\_\_\_\_\_ in the compound.
    - b. The ratio of their

← H<sub>2</sub>O---



- B. Chemical Bond
  - Strong attractive force between \_\_\_\_\_\_ or ions in a molecule or compound.
  - 2. Formed by:
    - electrons, (e-) by losing or gaining electrons.
    - electrons (e-)
- C. Stability
  - 1. Octet Rule
    - a. Most atoms form bonds in order to have \_\_\_\_\_\_\_ valence electrons (e-).
    - b. outer energy level.
    - c. Want to be like the Noble Gases (group 18).
  - Stability is the driving force behind bond formation. All atoms bond to become stable.
  - 3. Transferring e-



4. Sharing e-

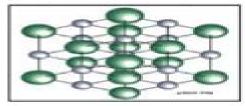


#### at. Kinds of Chemical Bonds

- A. Jonic Bond
  - Attraction between 2 charged ions.



- 2. Ions are \_\_\_\_\_\_atoms.
- A \_\_\_\_\_\_\_\_ is a \_\_\_\_\_\_ charged ion. Forms when an atom loses an electron. Formed by
- 4. A \_\_\_\_\_\_ is a \_\_\_\_\_ charged ion. Forms when an atom gains an electron. Formed by \_\_\_\_\_\_.
- Formed by electrons (e-) from a to a
- lons form a 3-D crystal lattice.



**Padhraic Smyth** 

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