

### Chemical Formula

Talking about proportions, molecules and relative of atoms present in a compound.

Molecular  $\text{C}_{10}\text{H}_{12}\text{O}_4$  (molecular)

Formula represents the number of atoms of each in a single molecule.

Empirical  $\text{C}_5\text{H}_6\text{O}_2$  (molecular)

Formula represents the simplest ratio of positive and negative ions in a compound.

### Atomic Mass

The weight of one mole of atoms of an element.

### Atomic

$12$  g atoms of carbon

$16$  g atoms of oxygen

### Molecular

$108$  g molecules

$160$  g molecules

$180$  g molecules

# Note Taking Guide Episode 601 Chemical Formula Answers

**D Siedentop**



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