Six Types of Chemical Reaction Worksheet Answers

Balance the following reactions and indicate which of the six types of chemical reaction are being represented:

a) 2 NaBr + 1 Ca(OH)₂ → 1 CaBr₂ + 2 NaOH

Type of reaction: double displacement

b) 2 NH₃+ 1 H₂SO₄ → 1 (NH₄)₂SO₄

Type of reaction: synthesis

c) 4 C₅H₉O + 27 O₂ → 20 CO₂ + 18 H₂O

Type of reaction: combustion

d) 3 Pb + 2 H₃PO₄ → 3 H₂ + 1 Pb₃(PO₄)₂

Type of reaction: single displacement

e) 1 Li₃N + 3 NH₄NO₃ → 3 LiNO₃ + 1 (NH₄)₃N

Type of reaction: double displacement

f) 3 HBr + 1 Al(OH)₃ → 3 H₂O + 1 AlBr₃

Type of reaction: acid-base (double displacement)

g) Na₃PO₄ + 3 KOH → 3 NaOH + K₃PO₄

Reaction Type __double displacement____

Reaction Type __double displacement____

h) MgCl₂ + Li₂CO₃ → MgCO₃ + 2 LiCl

Reaction Type ____complete combustion___

i) C₆H₁₂ + 9 O₂ → 6 CO₂ + 6 H₂O

Reaction Type ____single displacement ____

j) Pb + FeSO₄ → PbSO₄ + Fe

Reaction Type ____decomposition_____

k) CaCO₂ → CaO + CO₂

Reaction Type ___synthesis_____

I) P₄ + 3 O₂ → 2 P₂O₃

Reaction Type ___double displacement____

m) 2 RbNO₃ + BeF₂ → Be(NO₃)₂ + 2 RbF

Reaction Type single displacement

n) 2 AgNO₃ + Cu → Cu(NO₃)₂ + 2 Ag

Reaction Type ____complete combustion____

O₃H₆O + 4 O₂ → 3 CO₂ + 3 H₂O

Reaction Type ___synthesis____

p) 2 C₅H₅ + Fe → Fe(C₅H₅)₂

a) SeCl₀ + O₂ → SeO₂ + 3Cl₂

Reaction Type ___single displacement____

r) 2 Mgl₂ + Mn(SO₃)₂ → 2 MgSO₃ + Mnl₄

Reaction Type ____double displacement____

s) O₂ → O + O₂

Reaction Type ____decomposition_____

t) 2 NO₂ → 2 O₂ + N₂

Reaction Type _____decomposition_____

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Louis Agassiz Test, H. M. McLaughlin

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