

First Year Higher Secondary Examination

PART – III

CHEMISTRY

Maximum: 60 Scores

Time: 2 Hours

Cool-off time: 15 Minutes

1. (a) Determine the number of moles present in 0.55 mg of electrons.
 (i) 1 mole (ii) 2 mole (iii) 1.5 mole (iv) 0.5 mole (1)
 (b) Give the empirical formula of the following. (2)
 $C_8H_{12}O_6$, C_6H_6 , CH_3COOH , $C_6H_6Cl_6$
- (c) Two elements, carbon and hydrogen combine to form C_2H_6 , C_2H_4 and C_2H_2 , Identify the law illustrated here. (1)
2. (a) (i) Write the electronic configuration of chromium ($Z = 24$). (1)
 (ii) Find the number of electrons in the subshells with azimuthal quantum number $l = 2$. (1)
 (iii) Represent the orbital with quantum numbers $n = 1$ and $l = 0$ (1)
 (b) Give the mathematical representation of Heisenberg's uncertainty principle and its one important significance. (2)
3. Electron gain enthalpy is one of the important periodic property.
 (a) Define electron gain enthalpy. (1)
 (b) Explain any two factors affecting electron gain enthalpy. (2)
 (c) Write the oxidation state and co-valency of Al in $(AlF_6)^{3-}$. (1)
4. The geometry of the molecule is decided by type of hybridization.
 (a) Discuss the shape of PCl_5 molecules using hybridization. (2)
 (b) Give the reason for the high reactivity of PCl_5 . (2)
 (c) Isoelectric species have the same bond order. Among the following, choose the pair having same bond order.
 CN^- , $\bar{O}_2$, NO^+ , CN^+ . (1)

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Y Pai

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Table of Contents Polytechnic 1st Year Chemistry Question Paper

1. Understanding the eBook Polytechnic 1st Year Chemistry Question Paper
 - The Rise of Digital Reading Polytechnic 1st Year Chemistry Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Polytechnic 1st Year Chemistry Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Polytechnic 1st Year Chemistry Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Polytechnic 1st Year Chemistry Question Paper
 - Personalized Recommendations
 - Polytechnic 1st Year Chemistry Question Paper User Reviews and Ratings
 - Polytechnic 1st Year Chemistry Question Paper and Bestseller Lists
5. Accessing Polytechnic 1st Year Chemistry Question Paper Free and Paid eBooks
 - Polytechnic 1st Year Chemistry Question Paper Public Domain eBooks
 - Polytechnic 1st Year Chemistry Question Paper eBook Subscription Services
 - Polytechnic 1st Year Chemistry Question Paper Budget-Friendly Options
6. Navigating Polytechnic 1st Year Chemistry Question Paper eBook Formats

- ePub, PDF, MOBI, and More
- Polytechnic 1st Year Chemistry Question Paper Compatibility with Devices
- Polytechnic 1st Year Chemistry Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Polytechnic 1st Year Chemistry Question Paper
 - Highlighting and Note-Taking Polytechnic 1st Year Chemistry Question Paper
 - Interactive Elements Polytechnic 1st Year Chemistry Question Paper
- 8. Staying Engaged with Polytechnic 1st Year Chemistry Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Polytechnic 1st Year Chemistry Question Paper
- 9. Balancing eBooks and Physical Books Polytechnic 1st Year Chemistry Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Polytechnic 1st Year Chemistry Question Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Polytechnic 1st Year Chemistry Question Paper
 - Setting Reading Goals Polytechnic 1st Year Chemistry Question Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Polytechnic 1st Year Chemistry Question Paper
 - Fact-Checking eBook Content of Polytechnic 1st Year Chemistry Question Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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