

Name: \_\_\_\_\_

# CLASSIFYING CHEMICAL REACTIONS

Classify the following chemical reactions as one of the following: synthesis, decomposition, single-replacement, double-replacement, or combustion.

$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
$2\text{K} + 2\text{H}_2\text{O} \rightarrow \text{H}_2 + 2\text{KOH}$	$\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3$
$2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$	$\text{Ca} + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2$
$2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$	$\text{Pb(NO}_3)_2 + 2\text{HCl} \rightarrow \text{PbCl}_2 + 2\text{HNO}_3$
$\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$	$2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
$2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$	$\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
$\text{Pb(NO}_3)_2 + 2\text{K} \rightarrow \text{Pb} + 2\text{KNO}_3$	$\text{H}_2\text{CO}_3 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
$2\text{Fe} + \text{O}_2 \rightarrow 2\text{FeO}$	$\text{Zn} + 2\text{HCl} \rightarrow \text{H}_2 + \text{ZnCl}_2$
$\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$	$2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$

# Physical Science If8767 Classifying Chemical Reactions

**Michael Baer**



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## **Table of Contents Physical Science If8767 Classifying Chemical Reactions**

1. Understanding the eBook Physical Science If8767 Classifying Chemical Reactions
  - The Rise of Digital Reading Physical Science If8767 Classifying Chemical Reactions
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Classifying Chemical Reactions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science If8767 Classifying Chemical Reactions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Classifying Chemical Reactions
  - Personalized Recommendations
  - Physical Science If8767 Classifying Chemical Reactions User Reviews and Ratings

- Physical Science If8767 Classifying Chemical Reactions and Bestseller Lists
- 5. Accessing Physical Science If8767 Classifying Chemical Reactions Free and Paid eBooks
  - Physical Science If8767 Classifying Chemical Reactions Public Domain eBooks
  - Physical Science If8767 Classifying Chemical Reactions eBook Subscription Services
  - Physical Science If8767 Classifying Chemical Reactions Budget-Friendly Options
- 6. Navigating Physical Science If8767 Classifying Chemical Reactions eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science If8767 Classifying Chemical Reactions Compatibility with Devices
  - Physical Science If8767 Classifying Chemical Reactions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science If8767 Classifying Chemical Reactions
  - Highlighting and Note-Taking Physical Science If8767 Classifying Chemical Reactions
  - Interactive Elements Physical Science If8767 Classifying Chemical Reactions
- 8. Staying Engaged with Physical Science If8767 Classifying Chemical Reactions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science If8767 Classifying Chemical Reactions
- 9. Balancing eBooks and Physical Books Physical Science If8767 Classifying Chemical Reactions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science If8767 Classifying Chemical Reactions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science If8767 Classifying Chemical Reactions
  - Setting Reading Goals Physical Science If8767 Classifying Chemical Reactions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science If8767 Classifying Chemical Reactions
  - Fact-Checking eBook Content of Physical Science If8767 Classifying Chemical Reactions
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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