

- (b) A solution is prepared by the addition of 10.0 mL of 0.100 M CH_3COOH and 10.0 mL of 0.100 M NaOH .

- (c) The solution is a buffer solution of CH_3COOH and CH_3COO^- .

The total volume is 20.0 mL. The initial concentration of CH_3COOH is 0.100 M.

In this case, we should consider that the initial concentration of CH_3COOH is 0.100 M.

At the initial concentration of 0.100 M and 0.100 M:

$$[\text{CH}_3\text{COOH}] = 0.100 \text{ M} \quad [\text{CH}_3\text{COO}^-] = 0.100 \text{ M}$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad [\text{CH}_3\text{COO}^-] = \frac{[\text{CH}_3\text{COOH}]}{K_a} = \frac{0.100 \text{ M}}{1.8 \times 10^{-5}} = 5.56 \times 10^3 \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$\text{pH} = 4.74$$

Reaction: $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$
Initial: 0.100 M, 0.100 M, 0.000 M, 0.000 M
Change: -0.100 M, -0.100 M, +0.100 M, +0.100 M
Equilibrium: 0.000 M, 0.000 M, 0.100 M, 0.100 M
The solution is a buffer solution.

- (b) The solution is a buffer solution of CH_3COOH and CH_3COO^- . The initial concentration of CH_3COOH is 0.100 M.

In this case, we should consider that the initial concentration of CH_3COOH is 0.100 M.



- (c) The solution is a buffer solution of CH_3COOH and CH_3COO^- . The initial concentration of CH_3COOH is 0.100 M.

In this case, we should consider that the initial concentration of CH_3COOH is 0.100 M.

- (d) The solution is a buffer solution of CH_3COOH and CH_3COO^- .

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad [\text{CH}_3\text{COO}^-] = \frac{[\text{CH}_3\text{COOH}]}{K_a} = \frac{0.100 \text{ M}}{1.8 \times 10^{-5}} = 5.56 \times 10^3 \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$\text{pH} = 4.74$$

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Table of Contents Ocr Chemistry A Past Papers 23

1. Understanding the eBook Ocr Chemistry A Past Papers 23
 - The Rise of Digital Reading Ocr Chemistry A Past Papers 23
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry A Past Papers 23
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry A Past Papers 23
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry A Past Papers 23
 - Personalized Recommendations
 - Ocr Chemistry A Past Papers 23 User Reviews and Ratings
 - Ocr Chemistry A Past Papers 23 and Bestseller Lists
5. Accessing Ocr Chemistry A Past Papers 23 Free and Paid eBooks
 - Ocr Chemistry A Past Papers 23 Public Domain eBooks
 - Ocr Chemistry A Past Papers 23 eBook Subscription Services
 - Ocr Chemistry A Past Papers 23 Budget-Friendly Options
6. Navigating Ocr Chemistry A Past Papers 23 eBook Formats

- ePub, PDF, MOBI, and More
 - Ocr Chemistry A Past Papers 23 Compatibility with Devices
 - Ocr Chemistry A Past Papers 23 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry A Past Papers 23
 - Highlighting and Note-Taking Ocr Chemistry A Past Papers 23
 - Interactive Elements Ocr Chemistry A Past Papers 23
 8. Staying Engaged with Ocr Chemistry A Past Papers 23
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry A Past Papers 23
 9. Balancing eBooks and Physical Books Ocr Chemistry A Past Papers 23
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry A Past Papers 23
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ocr Chemistry A Past Papers 23
 - Setting Reading Goals Ocr Chemistry A Past Papers 23
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ocr Chemistry A Past Papers 23
 - Fact-Checking eBook Content of Ocr Chemistry A Past Papers 23
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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