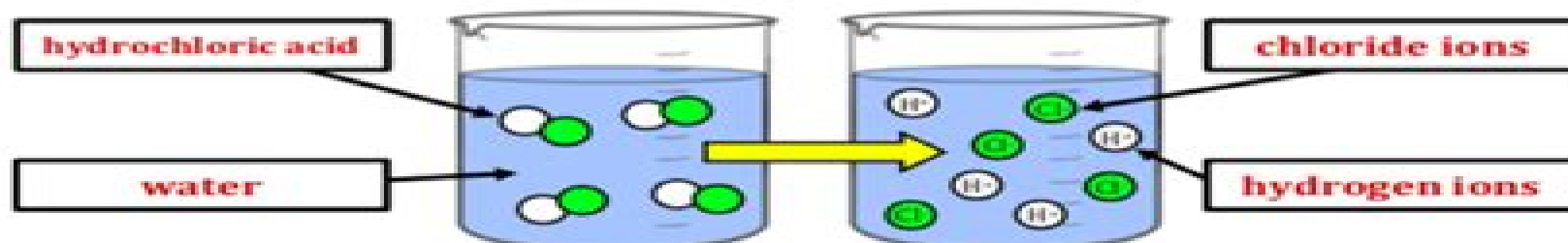


Introduction to Acids and Bases

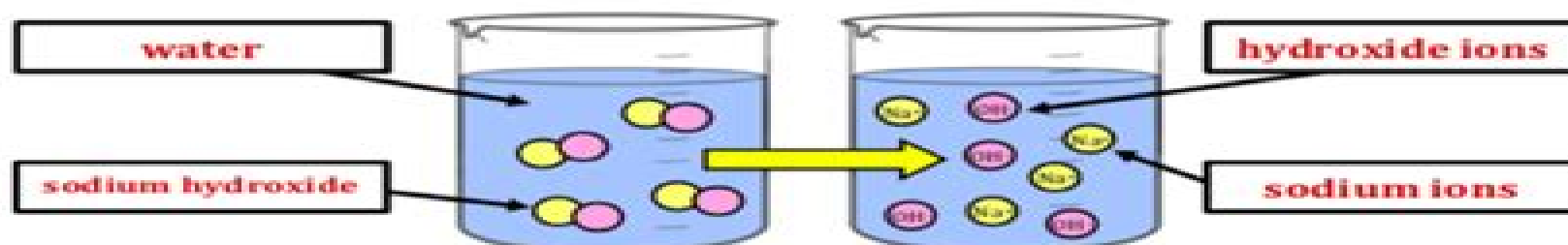
According to the **Arrhenius** definition:

- **Acids** are substances that release **hydrogen** ions in water (H^+)
- **Bases** are substances that release **hydroxide** ions in water (OH^-)

Instructions: The diagram below shows hydrochloric acid dissolving in water.
Label water, hydrochloric acid, hydrogen ions, and chloride ions.



Instructions: The diagram below shows sodium hydroxide dissolving in water.
Label water, sodium hydroxide, sodium ions, and hydroxide ions.



Instructions: Identify whether the chemicals below would be *acids* or *bases*.

- | | |
|----------------------------|-------------|
| 1) HBr | <u>acid</u> |
| 2) LiOH | <u>base</u> |
| 3) H_2SO_4 | <u>acid</u> |

- | | |
|-----------------------------|-------------|
| 4) $\text{Ca}(\text{OH})_2$ | <u>base</u> |
| 5) HNO_3 | <u>acid</u> |
| 6) KOH | <u>base</u> |

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Judd E. Hollander



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