

Chemical Formula

Talking about proportions molecules and relative of atoms present in a compound

Molecular $\text{C}_{10}\text{H}_{12}\text{O}_4$ (molecular)

Formula represents the number of atoms of each element in a single molecule

Empirical $\text{C}_5\text{H}_6\text{O}_2$ (molecular)

Formula represents the simplest ratio of positive and negative ions in a compound

Atomic Mass

The average mass of atoms present in a sample

Isotopes

^{12}C stable isotope carbon

^{13}C stable isotope carbon

Radioisotopes

^{14}C radioactive isotope carbon

^{15}O radioactive isotope oxygen

^{16}O stable isotope oxygen

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Siddappa N.Byrareddy



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